

M A buffer is made using the conjugate acid-base pair acetic acid and acetate ions. Write chemical equations showing the reactions that take place when H^+ and when OH^- are added to the buffer.

Materials

Chemical Analysis

Acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, 0.1 M, 20 ml.

Ammonia, NH₃, 0.1 M, 20 ml.Ammonium chloride, NH_4Cl , 0.1 M, 20 ml.

Hydrochloric acid, HCl, 0.1 M, 5 ml.

Phenylhydrazine, 1.0 g, 1 ml

Sodium acetate, NaCH_3CO_2 , 0.1 M, 20 ml.

Abstract

Balderson, C. 1991.

Features: 100- and 400-

Abstract—Cancer patients 3

Buret, ring stand, and beuret clamps

Eluent flow rate, 1.50 mL/min

Sodium chloride, NaCl, 0.1 M, 40 ml.

Sodium Hydroxide, NaOH, 0.1 M, 5 ml.

Sodium hydroxide, NaOH, 0.2 M, 30 ml.

Unknown solid acid crystals: 2.0 g

Water, distilled or deionized

$$(0.057M)(.100L) = 0.0057\text{mol}$$

$$[C_2H_3O_2^-] = \frac{K_a [HC_2H_3O_2]}{[H^+]} = \frac{(1.8 \times 10^{-5})(0.10)}{3.16 \times 10^{-5}}$$

100

Wash bottle with distilled or deionized water.

Flinn Buffer Solutions Lab Answer Key

S Ashworth

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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