

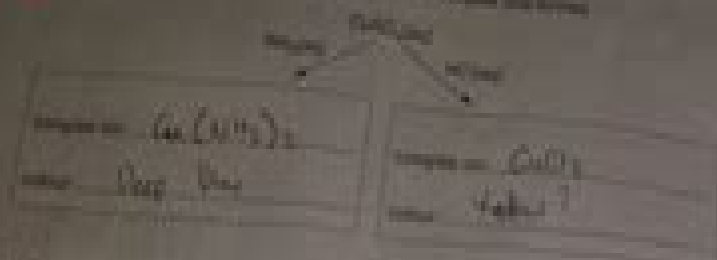
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a copper complex.

In the boxes, write the formula and name of the ligands and solvent.



(b) Cu^{2+} has two d-orbitals in d^9 with two unpaired electrons and two water molecules. The d-orbitals are in a tetrahedral field.

The ground state of the d-orbitals in d^9 is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term tetrahedral field?

A field where four ligands form tetrahedral arrangement around central metal ion.

(d) The complex is a copper(II) complex.

The reaction is about the synthesis of a copper complex. The reaction is shown below.



Complex A

Complete the 3D diagram of the other two geometries of A. You do not need to include any charges.

Indicate with ticks whether the stereoisomers are cis, trans, optical or a combination of these types.

Structure														
Type	<table border="1"> <tr> <td>cis</td> <td></td> </tr> <tr> <td>trans</td> <td>✓</td> </tr> <tr> <td>optical</td> <td></td> </tr> </table>	cis		trans	✓	optical		<table border="1"> <tr> <td>cis</td> <td>✓</td> </tr> <tr> <td>trans</td> <td></td> </tr> <tr> <td>optical</td> <td>✓</td> </tr> </table>	cis	✓	trans		optical	✓
cis														
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optical														
cis	✓													
trans														
optical	✓													

(e) What is the empirical formula, including the charge, of the complex in A?

$[\text{Cu}(\text{NH}_3)_2(\text{H}_2\text{O})_2]^{2+}$

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R Bogdan



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