

May 2009 TZ1 Question 1

- (b) For part of her extended essay investigation into the efficiency of the process, a student reacted a pure sample of a vegetable oil (where $R = C_{17}H_{33}$) with methanol. The raw data recorded for the reaction is below.

Mass of oil	= 1013.0 g
Mass of methanol	= 200.0 g
Mass of sodium hydroxide	= 3.5 g
Mass of biodiesel produced	= 811.0 g

The relative molecular mass of the oil used by the student is 885.6. Calculate the amount (in moles) of the oil and the methanol used, and hence the amount (in moles) of excess methanol.

[3]

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May 2009 TZ2 Question 1

- (b) A student reacted some salicylic acid with excess ethanoic anhydride. Impure solid aspirin was obtained by filtering the reaction mixture. Pure aspirin was obtained by recrystallization. The following table shows the data recorded by the student.

Mass of salicylic acid used	3.15 ± 0.02 g
Mass of pure aspirin obtained	2.50 ± 0.02 g

- (i) Determine the amount, in mol, of salicylic acid, $C_6H_4(OH)COOH$, used.

[2]

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- (ii) Calculate the theoretical yield, in g, of aspirin, $C_6H_4(OCOCH_3)COOH$.

[2]

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May 20chemistry Sl Paper

RD Boyd



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Table of Contents May 20chemistry Sl Paper

1. Understanding the eBook May 20chemistry Sl Paper
 - The Rise of Digital Reading May 20chemistry Sl Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying May 20chemistry Sl Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an May 20chemistry Sl Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from May 20chemistry Sl Paper
 - Personalized Recommendations
 - May 20chemistry Sl Paper User Reviews and Ratings
 - May 20chemistry Sl Paper and Bestseller Lists
5. Accessing May 20chemistry Sl Paper Free and Paid eBooks
 - May 20chemistry Sl Paper Public Domain eBooks
 - May 20chemistry Sl Paper eBook Subscription Services
 - May 20chemistry Sl Paper Budget-Friendly Options
6. Navigating May 20chemistry Sl Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - May 20chemistry Sl Paper Compatibility with Devices
 - May 20chemistry Sl Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of May 20chemistry Sl Paper
 - Highlighting and Note-Taking May 20chemistry Sl Paper
 - Interactive Elements May 20chemistry Sl Paper
8. Staying Engaged with May 20chemistry Sl Paper

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers May 20chemistry Sl Paper
9. Balancing eBooks and Physical Books May 20chemistry Sl Paper
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection May 20chemistry Sl Paper
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine May 20chemistry Sl Paper
- Setting Reading Goals May 20chemistry Sl Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of May 20chemistry Sl Paper
- Fact-Checking eBook Content of May 20chemistry Sl Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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