

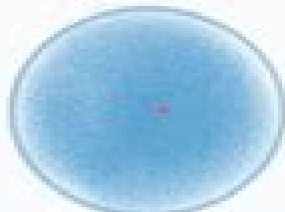
ELECTRON PROBABILITIES OF HYDROGEN ATOMS

HYDROGEN ATOM A



H
1s

HYDROGEN ATOM B



H
1s

FORMATION OF
COVALENT BOND
→

COVALENT BONDS

BONDING INTERACTIONS BETWEEN NONMETALS

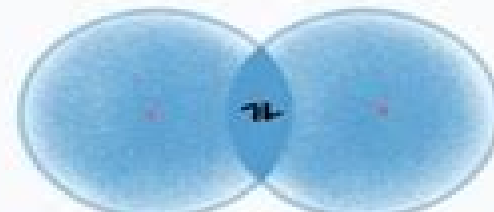
WHEN DISCUSSING IONIC BONDING, WE STUDIED HOW THE DONATION AND ACCEPTANCE OF ELECTRONS RESULTED IN AN ELECTROSTATIC INTERACTION RESULTING IN AN IONIC BOND BETWEEN METALS AND NONMETALS. HOWEVER, COVALENT BONDS ARE FORMED THROUGH AN ENTIRELY DIFFERENT METHOD - SHARING ELECTRONS. THIS SHARING OF ELECTRONS IS FORMED BETWEEN ATOMS WITH HIGH IONIZATION ENERGIES (NONMETALS), SO NEITHER CONTENDER WANTS TO LOSE ELECTRONS FROM THEIR VALENCE SHELL. INSTEAD, THE TWO ATOMS SHARE THEIR VALENCE ELECTRONS BY MERGING THEIR UNFILLED ORBITALS, PAIRING THE VALENCE ELECTRONS THROUGHOUT THE FORMATION OF THE BOND. THE TWO ATOMS ARE ATTRACTED TO ONE ANOTHER SINCE THE BONDING ELECTRONS SIT BETWEEN THE TWO NUCLEI - THIS ATTRACTION FORCE PULLS THE ATOMS CLOSER, LOWERING THE POTENTIAL ENERGY THROUGH THE FORMATION OF THE BOND.

THROUGH SHARING ELECTRONS, NEITHER ELECTRON HAS TO ABSORB ENERGY TO RELEASE AN ELECTRON, AND BOTH BONDING ATOMS ARE FURTHER STABILIZED BY FILLING THEIR ORBITALS (INCREASING THE SYMMETRY OF THEIR VALENCE SHELL).

FOR EXAMPLE, H_2 IS A MOLECULE FORMED FROM A COVALENT BOND OF TWO HYDROGEN ATOMS BY BONDING THE SINGLE UNPAIRED ELECTRONS FROM EACH ATOM (WITH OPPOSING SPIN STATES - PAULI EXCLUSION PRINCIPLE).

ELECTRON PROBABILITY OF HYDROGEN (H_2) MOLECULE

HYDROGEN (H_2) MOLECULE



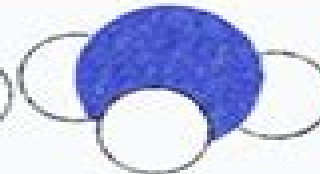
H:H

SHARING ELECTRON DENSITY:
INCREASING DENSITY BETWEEN ATOMS

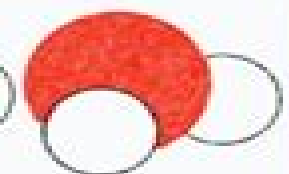
EXAMPLE OF COVALENT MOLECULES



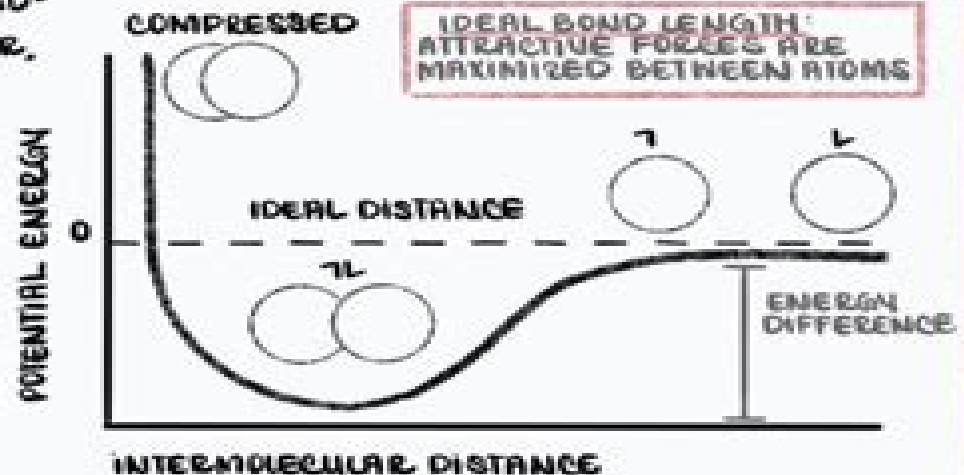
METHANE (CH_4)



AMMONIA (NH_3)



WATER (H_2O)



General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide

JG Myers



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1. Understanding the eBook General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - The Rise of Digital Reading General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - User-Friendly Interface

4. Exploring eBook Recommendations from General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Personalized Recommendations
 - General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide User Reviews and Ratings
 - General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide and Bestseller Lists
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 - General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide Compatibility with Devices
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7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Interactive Elements General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
8. Staying Engaged with General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
9. Balancing eBooks and Physical Books General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Setting Reading Goals General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Fact-Checking eBook Content of General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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