

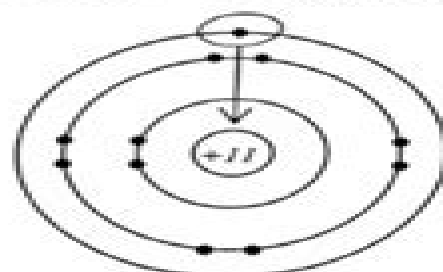
Advanced Periodic Trends

How do the particles in an atom interact to dictate periodic properties of elements?

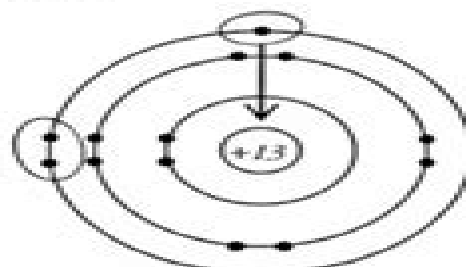
Why?

Previously you have learned about Coulombic attraction and how it governs the trends in properties among the elements of the Periodic Table. For example, as you move across a period, atoms have more protons in the nucleus, which pulls the electrons in tighter making smaller, more electronegative atoms. As you proceed down a column of the periodic table, the valence electrons get further away, and the atoms get larger and less electronegative. Inquisitive students will always ask, "If more protons in the nucleus pull electrons in tighter, then why don't atoms get smaller when going down a column?" That is a good question, and one that you will investigate in this activity.

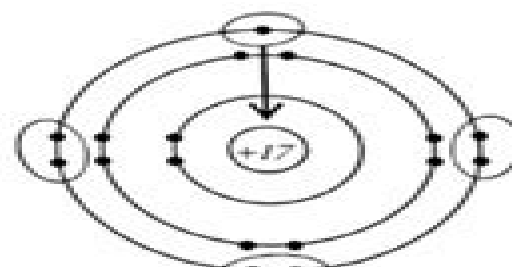
Model 1 – Period 3 Elements



Sodium



Aluminum



Chlorine

Atomic Number	11	13	17
Number of Electrons	11	13	17
Number of Valence Electrons	1	3	7
Core Charge	+1	+3	+7
Atomic Radius	186 pm	143 pm	99 pm
1st Ionization Energy	496 kJ/mol	578 kJ/mol	1251 kJ/mol
Electronegativity	0.9	1.5	3.0

1. Describe the relationship of the three elements in Model 1 with regard to their relative positions on the periodic table.

The three elements are in the same row of the periodic table.

2. Refer to a periodic table to complete the first three rows of the table in Model 1.

See Model 1.

Mini Pogil Discovering Periodic Trends

Brendan G. Carr



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Table of Contents Mini Pogil Discovering Periodic Trends

1. Understanding the eBook Mini Pogil Discovering Periodic Trends
 - The Rise of Digital Reading Mini Pogil Discovering Periodic Trends
 - Advantages of eBooks Over Traditional Books
2. Identifying Mini Pogil Discovering Periodic Trends
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mini Pogil Discovering Periodic Trends
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mini Pogil Discovering Periodic Trends
 - Personalized Recommendations
 - Mini Pogil Discovering Periodic Trends User Reviews and Ratings
 - Mini Pogil Discovering Periodic Trends and Bestseller Lists
5. Accessing Mini Pogil Discovering Periodic Trends Free and Paid eBooks
 - Mini Pogil Discovering Periodic Trends Public Domain eBooks
 - Mini Pogil Discovering Periodic Trends eBook Subscription Services
 - Mini Pogil Discovering Periodic Trends Budget-Friendly Options
6. Navigating Mini Pogil Discovering Periodic Trends eBook Formats

- ePub, PDF, MOBI, and More
- Mini Pogil Discovering Periodic Trends Compatibility with Devices
- Mini Pogil Discovering Periodic Trends Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mini Pogil Discovering Periodic Trends
 - Highlighting and Note-Taking Mini Pogil Discovering Periodic Trends
 - Interactive Elements Mini Pogil Discovering Periodic Trends
- 8. Staying Engaged with Mini Pogil Discovering Periodic Trends
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mini Pogil Discovering Periodic Trends
- 9. Balancing eBooks and Physical Books Mini Pogil Discovering Periodic Trends
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mini Pogil Discovering Periodic Trends
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mini Pogil Discovering Periodic Trends
 - Setting Reading Goals Mini Pogil Discovering Periodic Trends
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mini Pogil Discovering Periodic Trends
 - Fact-Checking eBook Content of Mini Pogil Discovering Periodic Trends
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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