

Orbitals and Bonding Test #14 SCH4U Name:

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Assessment of Learning

Multiple Choice: Circle the most correct answer. (15K)

1. After the 5s subshell of an atom is filled with electrons, the next electron added will enter the

- A) 4d subshell B) 5p subshell
C) 5d subshell D) 4f subshell

2. Which of the following is a correct set of quantum numbers for an electron in a 5f orbital?

- A) $n = 5, l = 3, m_l = +1$ B) $n = 5, l = 2, m_l = +3$ C) $n = 4, l = 3, m_l = 0$
D) $n = 4, l = 2, m_l = +1$ E) $n = 5, l = 4, m_l = 3$

3. Which one of the following combinations of quantum numbers is not allowed?

	n	l	m_l
A)	3	0	0
B)	3	1	-1
C)	3	2	2
D)	3	2	-1
E)	3	3	2

4. The orientation in space of an atomic orbital is associated with

- A) the principal quantum number (n). B) the angular momentum quantum number (l).
C) the magnetic quantum number (m_l). D) the spin quantum number (m_s).
E) none of the above

5. Consider the following electron configurations to answer the question that follows:

(i) $1s^2 2s^2 2p^5 3s^1$ (ii) $1s^2 2s^2 2p^4 3s^2$ (iii) $1s^2 2s^2 2p^5 3s^2 3p^1$

(iv) $1s^2 2s^2 2p^5 3s^2 3p^2$ (v) $1s^2 2s^2 2p^5 3s^2 3p^3$

The electron configuration belonging to the atom with the highest 1st ionization energy is

- A) (i) B) (ii) C) (iii) D) (iv) E) (v)

6. Which atomic orbitals overlap to form the C-O bond of dimethyl ether, $(CH_3)_2O$?

- A) C $2s^2$ + O $2s^2$ B) C $2s^2$ + O $2p$
C) C $2s^2$ + O $2s^2$ D) C $2s^2$ + O $2s^2$

General Chemistry Chapter 14 Covalent Bonding Orbitals Exam

LP Steffe

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape with a gradient, partially overlapping the bar's end.

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Table of Contents General Chemistry Chapter 14 Covalent Bonding Orbitals Exam

1. Understanding the eBook General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - The Rise of Digital Reading General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Personalized Recommendations
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exam User Reviews and Ratings
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exam and Bestseller Lists
5. Accessing General Chemistry Chapter 14 Covalent Bonding Orbitals Exam Free and Paid eBooks
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exam Public Domain eBooks
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exam eBook Subscription Services
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exam Budget-Friendly Options
6. Navigating General Chemistry Chapter 14 Covalent Bonding Orbitals Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exam Compatibility with Devices
 - General Chemistry Chapter 14 Covalent Bonding Orbitals Exam Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Highlighting and Note-Taking General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Interactive Elements General Chemistry Chapter 14 Covalent Bonding Orbitals Exam

8. Staying Engaged with General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
9. Balancing eBooks and Physical Books General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Setting Reading Goals General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Fact-Checking eBook Content of General Chemistry Chapter 14 Covalent Bonding Orbitals Exam
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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