

PRE-LAB QUESTIONS

1. Write names of the following monatomic ions.

Formula	Name
K^+	potassium ion
Ni^{2+}	nickel (II) ion
Zn^{2+}	zinc (II) ion
Cr^{3+}	chromium (III)
I^-	Iodide
S^{2-}	sulfide
P^{3-}	phosphide

2. Write the formulas of compounds formed by the following ions. For example, formula of compound with two
- Na^+
- ion and one
- O^{2-}
- ion is written as
- Na_2O
- . No charge should be written in the formulas of compounds.

a) compound with two Al^{3+} ions and three S^{2-} ions Al_2S_3 b) compound with one Ni^{2+} ions and two Cl^- ions $NiCl_2$

3. a) How many chloride ions (
- Cl^-
-) are needed to form a compound with a barium ion (
- Ba^{2+}
-)?
- 2

(Hint: A compound should have a net charge of zero. Therefore, sum of charges of ions in a compound should be zero.)

b) What is the formula of compound formed containing Ba^{2+} ion and Cl^- ion?
 Cl_2Ba

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W. Bassali



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Nuclear Chemistry [2] ,1968 Nuclear Chemistry L. Yaffe,2013-10-22 Nuclear Chemistry **An Introduction to Nuclear Chemistry** Glenn Theodore Seaborg,1946 Advanced Nuclear Chemistry S. K. Mehta,2010-01-01 *Nuclear Chemistry* R. K. Dave,2009-01-01 **Nuclear Chemistry Vol 1 & 2** L. Yaffe,1968 *Nuclear Chemistry* Axel N. Koskinen,2014-05-14 **Handbook of Nuclear Chemistry** Attila Vértes,Sándor Nagy,Zoltan Klencsár,Rezső G. Lovas,F. Rosch,2003 Principles of Nuclear Chemistry Peter A. C. McPherson,2017 **Nuclear Chemistry** Ritu Aneja,2004-01-01

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