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OPERATION MANUAL

First Edition

September 2000



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Tokyo, Japan

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FUJIFILM Corporation

Tokyo, Japan

# Fuji Fcr View Manual

**Sabine Zange**



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web 1 practice problem balance the redox reaction in basic solution  $\text{NO}_2^- + \text{Al} + \text{S} + \text{NH}_3 \rightarrow \text{AlO}_2^- + \text{N}_2 + \text{S}^{2-}$  practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of  $\text{OH}^-$   $\text{Br}_2 + \text{I}^- \rightarrow \text{Br}^- + \text{I}_2$  previous topic next topic

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web google classroom for the following redox equation  $\text{MnO}_4^- + \text{SO}_2 + 2\text{H}^+ \rightarrow \text{Mn}^{2+} + \text{H}_2\text{SO}_4$  what will be the stoichiometric coefficient of  $\text{H}_2\text{SO}_4$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow \text{Au}(\text{CN})_2^- + 2\text{e}^-$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- \rightarrow 2\text{Au}(\text{CN})_2^- + 2\text{H}_2\text{O} + \text{O}_2$  multiplied by a factor of 2  $2\text{H}_2\text{O} + \text{O}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{OH}^-$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh ions or the oh h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo and cr oh<sub>4</sub> in basic solution created by jay

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web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution  $4\text{Fe} + 6\text{H}_2\text{O} + 7\text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{NO} + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox

reaction in acid solution using the half reaction method 8 h<sub>2</sub>o<sub>2</sub> cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> 2 9 teo<sub>3</sub> 2 n<sub>2</sub>o<sub>4</sub> o<sub>2</sub> cr<sub>3</sub> te 10 reo<sub>4</sub> 11 pbo<sub>2</sub> io i<sub>2</sub> pb<sub>2</sub> 12 as no<sub>3</sub> io<sub>3</sub> re io<sub>3</sub> h<sub>2</sub>aso<sub>4</sub>

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h<sup>+</sup> 2no<sup>2-</sup> 2 c 6 h 6 o 6 2h<sup>+</sup> 2 o<sup>2-</sup> 2no choose 1 answer c 6 h 8 o 6 and no<sup>2-</sup> a c 6 h 8 o 6 and no<sup>2-</sup> no<sup>2-</sup> 2 and c 6 h 8 o 6 b no<sup>2-</sup> 2 and c 6 h 8 o 6 h<sup>+</sup> and no<sup>2-</sup> c h<sup>+</sup> and

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