

Atomic Structure Answer Key

1. A negatively-charged particle moving around the nucleus of an atom is a(n) _____.
 - a. compound
 - b. electron**
 - c. atom
3. Why does an atom have no net electric charge?
 - a. Its subatomic particles carry no electrical charge.
 - b. The positively charged protons cancel out the negatively charged neutrons.
 - c. The positively charged neutrons cancel out the negatively charged electrons.
 - d. The positively charged protons cancel out the negatively charged electrons.**
5. Compared to the charge and mass of a proton, an electron has
 - a. the same charge and a smaller mass.
 - b. the same charge and the same mass.
 - c. opposite charge and smaller mass.**
 - d. opposite charge and the same mass.
7. The electron cloud is a visual model of the probable locations of electrons in an atom.
 - a. True**
 - b. False
9. In Bohr-Rutherford diagrams, neutrons are typically located in the rings around the nucleus.
 - a. True
 - b. False**
2. A positively-charged particle in the nucleus of an atom is a(n) _____.
 - a. electron
 - b. proton**
 - c. element
4. What is an isotope?
 - a. An element with equal number of protons and neutrons
 - b. An elements with differing numbers of protons and electrons
 - c. An element with differing number of protons and neutrons**
 - d. An element with equal number of protons and electrons
6. A proton has approximately the same mass as _____.
 - a. a neutron**
 - b. a beta particle
 - c. an alpha particle
 - d. an electron
8. What is the nucleus made of?
 - a. protons
 - b. electrons
 - c. neutrons
 - d. protons and electrons
 - e. protons and neutrons**
 - f. electrons and neutrons
10. If atoms of a halogen nonmetal (Group 17) gain one electron, the atoms then have _____.
 - a. no valence electrons
 - b. 7 valence electrons
 - c. 8 valence electrons**
 - d. 17 valence electrons

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Clemens Wendtner



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