



PHYSICAL ANALYTICAL CHEMISTRY

Name _____ Section _____ Date _____

CHAPTER 6 WORKSHEET

1. Balance the following reactions in both acidic and basic media



To balance this, we should separate this into two half equations which are $\text{H}_2\text{PO}_2^- \rightarrow \text{H}_2\text{P}$ and $\text{H}_2 \rightarrow \text{H}^+$ which are balanced below:

balancing the reduction reaction



balancing the oxidation reaction



In combining the two half equations, make sure that the electrons are equal in each half equation.



To make this in basic medium, we add enough OH^- equals to H^+ .



To balance this, we should separate this into two half equations which are $\text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+}$ and $\text{I}^- \rightarrow \text{I}_2$ which are balanced below:

balancing the reduction reaction



balancing the oxidation reaction



Episode 1103 Chemistry Worksheet

Michael Rajnik



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