

## Physical Science Standards:

**SPS1:** Students will investigate our current understanding of the atom.

### a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
  - Proton has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Protons are also the elements' \_\_\_\_\_ number of the identity of the element, so protons are found by looking at the atomic number
  - Neutron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Neutrons are found by subtracting the \_\_\_\_\_ Number from the \_\_\_\_\_ Mass or the BIG #- SMALL #.
  - Electron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Electrons are equal to the number of \_\_\_\_\_ in a neutral atom.
  - The nucleus is made up of these 2 particles, \_\_\_\_\_ and \_\_\_\_\_.
  - The electrons are located in this region of the atom, known as the \_\_\_\_\_.
  - The electron cloud is divided into energy levels or energy shells, the first level will hold \_\_\_\_\_ electrons and the second level will hold \_\_\_\_\_ electrons.
  - The valence shell or valence level is the \_\_\_\_\_ energy shell of the atom.
- atomic mass and atomic number.
  - Atomic number is equal to the number of \_\_\_\_\_ in the atom and is also the atom's identity—it's the small number of the periodic table.
  - Atomic mass is equal to the average mass of all the isotopes of a particular element—it's the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
  - Define Isotope \_\_\_\_\_.
  - Ex. Carbon-12 and Carbon-14
    - Carbon 12 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
    - Carbon 14 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
- explain the relationship of the proton number to the element's identity.
  - The \_\_\_\_\_ number is equal to the number of \_\_\_\_\_ in the atom which is how the element is identified.
  - Ex. Nitrogen has an atom number of \_\_\_\_\_, which means that it has \_\_\_\_\_ protons.

molybdenum	← element name
42	← atomic number number of protons (Z)
Mo	← atomic symbol
95.94	← atomic mass A (this is an average mass)



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

# Eoct Physics Study Guide

**Rosina Ehmman**



## **Eoct Physics Study Guide:**

## **Eoct Physics Study Guide** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has be much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Eoct Physics Study Guide**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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### **Table of Contents Eoct Physics Study Guide**

1. Understanding the eBook Eoct Physics Study Guide
  - The Rise of Digital Reading Eoct Physics Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Eoct Physics Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Eoct Physics Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Eoct Physics Study Guide
  - Personalized Recommendations
  - Eoct Physics Study Guide User Reviews and Ratings
  - Eoct Physics Study Guide and Bestseller Lists

5. Accessing Eoct Physics Study Guide Free and Paid eBooks
  - Eoct Physics Study Guide Public Domain eBooks
  - Eoct Physics Study Guide eBook Subscription Services
  - Eoct Physics Study Guide Budget-Friendly Options
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  - Eoct Physics Study Guide Compatibility with Devices
  - Eoct Physics Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Eoct Physics Study Guide
  - Highlighting and Note-Taking Eoct Physics Study Guide
  - Interactive Elements Eoct Physics Study Guide
8. Staying Engaged with Eoct Physics Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Eoct Physics Study Guide
9. Balancing eBooks and Physical Books Eoct Physics Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Eoct Physics Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Eoct Physics Study Guide
  - Setting Reading Goals Eoct Physics Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Eoct Physics Study Guide
  - Fact-Checking eBook Content of Eoct Physics Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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