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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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configuration 3s² explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain hund s rule

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web the electron removed during the ionization of beryllium he 2s² is an s electron whereas the electron removed during the ionization of boron he 2s² 2p¹ is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s² ne 10 neon 1s² 2s² 2p⁶

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is 1s² 2s² 2p⁶ 3s² 3p⁶ 4s¹ or simply put as 4s¹ q2 what is the total number of electrons that can be filled in s p and d subshells

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web 2 locate ducts against wall or maximum of 51 mm away from wall 3 each wall anchor shall satisfy the following criteria unless other analysis is made a tensile load $3.8 \times$ duct weight safety factor 4 b shear load $1.2 \times$ duct weight safety factor 4 figure 4 8m support from wall hvac duct construction standards metal and flexible

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