

Question 13
Acids and Bases IV
Answered

13. Refer to Fig. 10.10 of Part I as added to a 0.1 M sodium acetate solution (0.4 M). What is the concentration of H^+ ions? Assume that the total H^+ density is constant. Assume that the overall volume of the solution does not change.

$$[\text{HAc}]_{\text{eq}} \text{ and } [\text{Ac}^-]_{\text{eq}} = 0.4 \text{ M}$$

$$\text{pH} = 4.75 \Rightarrow \log_{10} K_a = 4.75$$

Step 1: with iterative complexity. The addition of H^+ ions will shift the equilibrium of the dissociation of the base, according to Le Chatelier's principle. Assume H^+ will combine with Ac^- to form the protons HAc . This will continue until the product of $[\text{H}^+][\text{Ac}^-]$ equals K_a (iterative). This will decrease the concentration of H^+ in the solution.

14. A buffer solution is created by dissolving 10 mmol of sodium acetate. What is the pH of a 0.1 M solution of H_2A ($\text{p}K_1 = 4.2$) in H_2O ? Assume that the volume of the solution does not change.

- a. What the pH of the solution

14. Answer: 4.20
 14.100 mol H_2A + 0.100 mol Ac^- $\text{pH} = \text{pH}_0 + \log_{10} \frac{[\text{Ac}^-]}{[\text{HAc}]}$ $= -\log(10) + 0.100 + \log_{10} \frac{0.100}{0.100} = 4.2$
 14.100 mol

15. What the pH after 0.1 M and 0.1 M H_2A is added to the solution. Assume that the volume does not change.

15.100 mol H_2A + 0.100 mol Ac^- and 0.1 M H_2A
 15.100 mol

	H^+	\rightleftharpoons	Ac^-	\rightleftharpoons	pH
1	0.100 mol		0.100 mol		0.100 mol
2	0.100 mol		0.100 mol		0.100 mol
3	0.100 mol		0.100 mol		0.100 mol

15.100 mol $\frac{[\text{Ac}^-]}{[\text{HAc}]}$ $= 0.100 + 0.100$ $\text{pH} = 4.2$ 15.100 mol $\frac{[\text{Ac}^-]}{[\text{HAc}]}$ $= 0.100 + 0.100$ $\text{pH} = 4.2$

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Bruce A. Chabner, Dan L. Longo



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