



CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CH_4 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, resulting in oppositely charged ions. Example: $NaCl$, KBr .
- Metallic Bond:** Formed by the delocalization of valence electrons in a metal lattice. Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonds

Covalent bonds are formed by sharing electrons. They are typically found in nonmetals and their compounds. Ionic bonds are formed by the transfer of electrons. They are typically found in metals and nonmetals. Metallic bonds are formed by the delocalization of electrons. They are typically found in metals.

Science Real

IONIC BONDING

Atoms combine to form ionic compounds by transferring electrons. This process is called ionic bonding. The resulting ionic compound has properties different from the individual atoms.

Types of Ionic Bonds:

- Single Bond:** Formed by the transfer of one electron from one atom to another.
- Double Bond:** Formed by the transfer of two electrons from one atom to another.
- Triple Bond:** Formed by the transfer of three electrons from one atom to another.

Covalent Bonding

Atoms combine to form covalent compounds by sharing electrons. This process is called covalent bonding. The resulting covalent compound has properties different from the individual atoms.

Types of Covalent Bonds:

- Single Bond:** Formed by the sharing of one electron pair between two atoms.
- Double Bond:** Formed by the sharing of two electron pairs between two atoms.
- Triple Bond:** Formed by the sharing of three electron pairs between two atoms.

Metallic Bonding

Atoms combine to form metallic compounds by the delocalization of electrons. This process is called metallic bonding. The resulting metallic compound has properties different from the individual atoms.

Types of Metallic Bonds:

- Single Bond:** Formed by the delocalization of one electron from one atom to another.
- Double Bond:** Formed by the delocalization of two electrons from one atom to another.
- Triple Bond:** Formed by the delocalization of three electrons from one atom to another.

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