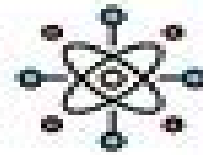


# CHEMICAL BONDS



## IONIC, COVALENT, & METALLIC

**FORMS OF CHEMICAL BONDS**  
Bonding: combination

1. **IONIC BONDING** occurs between a metal and a nonmetal. The metal atom loses one or more electrons to become a positively charged cation, and the nonmetal atom gains those electrons to become a negatively charged anion. The resulting ions are held together by a strong electrostatic attraction.



2. **COVALENT BONDING** occurs between two nonmetals. The atoms share one or more pairs of electrons, forming a shared electron pair that holds the atoms together.

3. **METALLIC BONDING** occurs between metal atoms. The atoms are held together by a sea of delocalized electrons that are shared among all the atoms in the metal lattice.

**CHEMICAL BONDS AND ELECTRONS**

Electrons are arranged in shells around the nucleus of an atom. The outermost shell is called the valence shell. The number of valence electrons determines the chemical properties of an element.

Atoms with one, two, or three valence electrons tend to lose electrons and form cations. Atoms with four, five, six, or seven valence electrons tend to share electrons and form covalent bonds. Atoms with eight valence electrons are stable and do not form bonds.



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**IONIC BONDING**

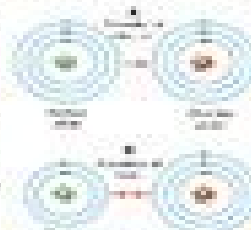

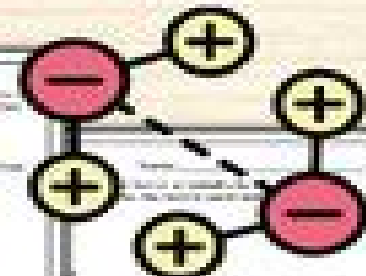
When a metal atom loses an electron, it becomes a positively charged cation. When a nonmetal atom gains an electron, it becomes a negatively charged anion. The resulting ions are held together by a strong electrostatic attraction, forming an ionic bond.

**COVALENT BONDING**

When two nonmetal atoms share one or more pairs of electrons, they form a covalent bond. The shared electrons are held between the two atoms, pulling them together.

**METALLIC BONDING**

When metal atoms are packed closely together, their valence electrons are shared among all the atoms, forming a sea of delocalized electrons. This sea of electrons holds the metal atoms together.



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**Xiang Xie**



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