

Thus, the oxidation number of sodium in the compound is $1+$ because one electron was transferred from the sodium atom. Because an electron is transferred to the chlorine atom, its oxidation number is $1-$.


- 7- In the chemical formula for any ionic compound, the symbol of the cation is always written first, followed by the symbol of the anion.
- 8- Subscripts, which are small numbers to the lower right of a symbol, represent the number of ions of each element in an ionic compound. If no subscript is written, it is assumed to be one.

Practice Problems **Chlorine** **Write Formulas**

Write formulas for the ionic compounds formed by the following ions.

19. potassium and iodide	21. aluminum and bromide
20. magnesium and chloride	22. sodium and nitride

23. **Challenge:** Write the general formula for the ionic compound formed by elements from the two groups shown on the periodic table at the right.



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Formulas for polyatomic ionic compounds:

- 9- polyatomic ions are ions made up of more than one atom. A polyatomic ion acts as an individual ion. Thus, the formula for a polyatomic compound follows the same rules used for a binary compound. Because a polyatomic ion exists as a unit, never change subscripts of the atoms within the ion.

Ionic Compounds Study Guide

Lei Shi

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Ionic Compounds Study Guide:

Uncover the mysteries within Crafted by is enigmatic creation, Discover the Intrigue in **Ionic Compounds Study Guide** . This downloadable ebook, shrouded in suspense, is available in a PDF format (*). Dive into a world of uncertainty and anticipation. Download now to unravel the secrets hidden within the pages.

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Table of Contents Ionic Compounds Study Guide

1. Understanding the eBook Ionic Compounds Study Guide
 - The Rise of Digital Reading Ionic Compounds Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Compounds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Compounds Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Compounds Study Guide
 - Personalized Recommendations
 - Ionic Compounds Study Guide User Reviews and Ratings
 - Ionic Compounds Study Guide and Bestseller Lists
5. Accessing Ionic Compounds Study Guide Free and Paid eBooks
 - Ionic Compounds Study Guide Public Domain eBooks
 - Ionic Compounds Study Guide eBook Subscription Services
 - Ionic Compounds Study Guide Budget-Friendly Options
6. Navigating Ionic Compounds Study Guide eBook Formats

- ePub, PDF, MOBI, and More
- Ionic Compounds Study Guide Compatibility with Devices
- Ionic Compounds Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Compounds Study Guide
 - Highlighting and Note-Taking Ionic Compounds Study Guide
 - Interactive Elements Ionic Compounds Study Guide
- 8. Staying Engaged with Ionic Compounds Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Compounds Study Guide
- 9. Balancing eBooks and Physical Books Ionic Compounds Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Compounds Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Compounds Study Guide
 - Setting Reading Goals Ionic Compounds Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Compounds Study Guide
 - Fact-Checking eBook Content of Ionic Compounds Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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