

Name Solutions

Date \_\_\_\_\_ Period \_\_\_\_\_

### Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers 1ve <sup>-</sup> to F	Li <sup>+</sup>
F	7 ve	accepts 1ve <sup>-</sup> from Li	F <sup>-</sup>

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$ $\text{Li} + \cdot\ddot{\text{Cl}}\cdot \rightarrow [\text{Li}]^+ [\ddot{\text{Cl}}:]^{-1}$	Li	1 ve	donate 1e <sup>-</sup>	Li <sup>+</sup>
	Cl	1 ve	accepts 1e <sup>-</sup>	Cl <sup>-</sup>
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ $\text{Ca} + \cdot\ddot{\text{O}}\cdot \rightarrow [\text{Ca}]^{+2} [\ddot{\text{O}}:]^{-2}$	Ca	2 ve	donates 2e <sup>-</sup>	Ca <sup>+2</sup>
	O	6 ve	accepts 2e <sup>-</sup>	O <sup>-2</sup>
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ $\text{Be} + \cdot\ddot{\text{F}}\cdot \rightarrow [\text{Be}]^{+2} [\ddot{\text{F}}:]^{-1}$	Be	2 ve	donates 2e <sup>-</sup> to each F	Be <sup>+2</sup>
	F	7 ve	each F accepts 1ve each	F <sup>-</sup>
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ $\text{Mg} + \cdot\ddot{\text{S}}\cdot \rightarrow [\text{Mg}]^{+2} [\ddot{\text{S}}:]^{-2}$	Mg	2 ve	donates 2e <sup>-</sup> to S	Mg <sup>+2</sup>
	S	6 ve	accepts 2e <sup>-</sup>	S <sup>-2</sup>
5) $\text{K} + \text{F} \Rightarrow \text{KF}$ $\text{K} + \cdot\ddot{\text{F}}\cdot \rightarrow [\text{K}]^+ [\ddot{\text{F}}:]^{-1}$	K	1 ve	donates 1ve <sup>-</sup>	K <sup>+1</sup>
	F	7 ve	accepts 1ve <sup>-</sup>	F <sup>-</sup>

# Ionic Bonds Review And Reinforce Worksheet Answers

**Siddappa N.Byrareddy**



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