# pH Properties of Buffer Solutions

#### Flinn AP Chemistry Laboratory #19

#### Introduction

One of the most important applications of acids and bases in chemistry and biology is that of buffers. A buffer solution resists rapid changes in pH when acids and bases are added to it. Every living cell contains natural buffer systems to maintain the constant pH needed for proper cell function. Many consumer products are also buffered to safeguard their activity. What are buffers made of? How do buffers maintain the delicate pH balance needed for life and health?

#### Concepts

- Buffers
- Weak acid-weak base
- · Conjugate acid-base pairs
- Dissociation constant
- Ideal buffer
- Background

The ability of buffers to resist changes in pH when acid or base is added is a result of their chemical composition. All buffers contain a mixture of a conjugate acid-base pair; either a weak acid (HA) and its conjugate base (A-), or a weak base (B), and its conjugate acid (BH+). Weak acids and bases both dissociate slightly in water (Reactions 1 and 2).

HA(aq) + H <sub>2</sub> O(I)	F→ H <sub>1</sub> O+(aq) + A-(aq)	Reaction 1
weak acid	conjugate base	
$B(aq) + H_1O(l) \leftarrow$	→ BH+(aq) + OH-(aq)	Reaction 2
www.k.hase	conjugate acid	

These reactions are reversible and both the weak acid and its conjugate base or the weak base and its conjugate acid are present in solution.

The equilibrium constant expression for these dissociation reactions are:

$$Ka = [H_3O+][A-]/[HA]$$
 Equation 1  
 $Kb = [OH-][BH+]/[B]$  Equation 2

Buffers control pH because the two buffering compounds, either HA and A- or B and BH+, are able neutralize both acids and bases added to the solution.

$$HA(aq) + OH - (aq) \rightarrow H_2O(1) + A - (aq)$$
 Reaction 3

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**George G. Manov** 

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buffer capacity which would make them acceptable as buffer materials In the past thirty years the concept of the term pH underwent a change from that of the simple definition of pH logCH to the more modern definition of pH logaH 1 when it was realized that the potential of the hydrogen electrode was a measure of the activity of the hydrogen ion rather than of its concentration Paradoxically however the activity of single ions cannot be determined from experiment without the use of some simplifying assumptions it is therefore impossible to establish a pH scale on a rigorous thermodynamic basis This point will be discussed later The problem of purification of salts is one primarily of analytical chemistry and the materials can be tested to determine their relative purity and reproducibility by any self consistent method Crystallization-related PH Changes During Freezing of Sodium Phosphate Buffer Solutions Gerardo Gómez, 1995 Measurement of PH Richard E. Lawn, Florence Elizabeth Prichard, 2003 Whether pH is being used to test a sample against a legal requirement or specification as part of an analytical method for monitoring and controlling a reaction as a process control in the chemical industry or for the environmental monitoring of waste and effluents it is important that all pH measurements are carried out in a logical and consistent manner paying careful attention to experimental procedures in order to obtain reliable results This guide provides scientists with the knowledge of how to do just that first by outlining the principles of pH measurement and buffer solutions pH meters and electrodes are then discussed including selection criteria and the care of electrodes Finally sections on making pH measurements and uncertainty are followed by a set of practical exercises Measurement of pH is one of the Practical Laboratory Skills Training Guides a series that aims to make achieving best practice easy These invaluable manuals will enable both experienced and inexperienced staff to get the essential basics of any experiment right simply by following the clear and easy to use instructions provided The guides are written by experienced scientists and include minimal theory plenty of practical exercises in order to assess competence and trouble shooting information Other titles are Measurement of Mass Measurement of Volume High Performance Liquid Chromatography and Gas Chromatography

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