

Empirical Formula and Molecular Formula

Lab 10: Molecular Formulas and Empirical Formulas

Lab 10

Students will determine the empirical formula of a compound by determining the mass percent composition of compounds. They will then use the empirical formula to determine the molecular formula of a compound. This activity will require an empirical formula.

Model 1: Comparison of Percent Composition and Empirical Formula

Compound	Empirical Formula	Molecular Formula	Empirical Formula	Percent Composition (Carbon)	Percent Composition (Hydrogen)
1	ethane	C_2H_6	CH_3		
2	propane	C_3H_8	CH_2		
3	ethylene (acetylene)	C_2H_2	CH		
4	butane	C_4H_{10}	C_2H_5		

1. Looking at the table, what are some similarities between the percent composition and empirical formula of Model 1. What are some differences? (You may calculate) below:
2. Identify the compounds in Model 1 that have the same empirical formula.
3. Identify the compounds in Model 1 that have the same percent composition.
4. Using the compounds in Model 1, suggest an explanation as to why compounds with different molecular formulas but the same empirical formula have the same percent composition.
5. For each substance in Model 1, give evidence that the molecular formula and empirical formula are not the same. (What mathematical operation did you use to determine the molecular formula? Is the molecular formula the same as the empirical formula?)

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