#### Empirical and Molecular Formula Practice KEY

#### Empirical Formula

36.5% Na, 25.4% S, and 38.1% O forms this compound.

Na - 36.5/23 = 1.56/0.8 = 2

S - 25.4/32 = 0.8/0.8 = 1

0 - 38.1/16 = 2.4/0.8 = 3

Na<sub>2</sub>5O<sub>3</sub>

A compound made of 72.2% Magnesium and 27.8% nitrogen by mass.

Mg - 72.2/24.3 = 3/2 = 1.5(2) = 3

N - 27.8/14 = 2/2 = 1(2) = 2

MigaNia

A chromium compound with 68.4% Cr and 31.6% O.

Cr - 68.4/52 = 1.315/1.315 = 1(2) = 2

O - 31.6/16 = 1.975/1.315 = 1.5(2) = 3

Cr<sub>2</sub>O<sub>3</sub>

4. A compound with the composition of 24% C, 7% H, 38% F, and 31% P.

C - 24/12 = 2/1 = 2

H - 7/1 = 7/1 = 7

F - 38/19 = 2/1 = 2

P - 31/31 = 1/1 = 1

C<sub>2</sub>H<sub>2</sub>F<sub>2</sub>P

5. 63.5% silver, 8.2% nitrogen, and 28.3% oxygen make up this compound.

Ag - 63.5/107.87 = 0.589/0.585 = 1

N - 8.2/14.01 = 0.585/0.585 = 1

0 - 28.3/16 = 1.769/0.585 = 3

AgNO<sub>1</sub>

8.81g C and 91.2g Chlorine forms this compound.

C - 8.81/12 = 0.73/0.73 = 1(2) = 2

CI - 91.2/35 = 2.61/0.73 = 3.5(2) = 7

C<sub>2</sub>Cl<sub>2</sub>

7. A compound in a soft drink that has 0.31g hydrogen, 3.2g phosphorus, and 6.5g oxygen.

H - 0.31/1 = 0.31/.1 = 3.1

P - 3.2/31 = 0.10/.1 = 1

0 - 6.5/16 = 0.40/.1 = 4

HIPO.

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**DP Hallahan** 

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