

Name: _____
Period: _____

Fall Semester Review – IPC

IPC Chemistry Review

Fill in the table at the right from your periodic table.

Element	Chemical Symbol	Atomic #	Valence Electrons	Oxidation #	# of Protons	Atomic Mass	Metal or Nonmetal	Group or Luster Electrons	# of Full Electron Levels
Oxygen									
Argon									
Calcium									
Nitrogen									
Potassium									
Sulfur									
Helium									

Find the molar mass of the following compounds.

Calcium oxide (CaO)

Water

Classify (ionic, covalent or polyatomic) and name the following compounds.

1. CO₂ _____

2. Na₂O _____

3. BaCrO₄ _____

Draw the Lewis Dot Diagrams for

1. Carbon

2. Sodium

Make balanced ionic compounds from:

1. Li + O _____

3. Be + NO₃⁻ _____

2. Na + Ar _____

4. Ca + O _____

With dot diagrams draw the covalent bond of O₂

With dot diagrams the covalent bond of CO₂

Matching:

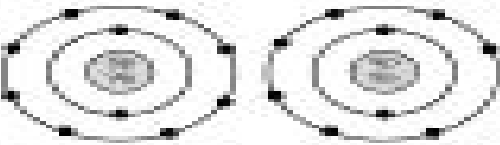
- A. Mixture
- B. Compound
- C. Element
- D. Mixture
- E. Heterogeneous
- F. Homogeneous

- _____ A substance made up of two or more atoms that must be separated by chemical means.
- _____ Something made up of 2 or more different atoms that can be separated by physical means.
- _____ A substance that is pure and made up of only one type of atom.
- _____ Anything that has mass and takes up space.
- _____ A mixture that is the same throughout.
- _____ A mixture that is different throughout.

1. Valence electrons are the outermost electrons of an atom that are involved in chemical bonding? True/False

2. Lithium has _____ valence electrons, will (lose or gain) _____ electrons and become (positive or negative). This is why lithium's oxidation number is _____.

3. A metal and non-metal will form a _____ compound. Two non-metals form a _____ compound.



Which one is the ion?

Give the ion notation for the ion.

Atom A

Atom B

Matching:

- A. Ion
- B. Element
- C. Isotope
- D. Oil

- _____ If you change the number of protons, you change the _____.
- _____ If you change the number of neutrons, you change the _____.
- _____ If you change the number of electrons, you change the _____.
- _____ Every 3,000 miles you should change a car's _____.

Matching:

- A. Law of Conservation of Mass
- B. Matter
- C. Centimeter
- D. Liter
- E. Gram

- _____ Unit of length: equal to 3.2 ft.
- _____ Mass is neither destroyed or created in chemical reactions.
- _____ Unit of mass: about 1 dollar bill.
- _____ 1/100th of mass: width of pinky finger.
- _____ Unit of volume: just bigger than a quart.

Matching:

- A. Solid
- B. Liquid
- C. Gas
- D. Oil

- _____ Molecules that are tightly packed and retain their shape and size.
- _____ Molecules that bounce off of each other, can be compressed, and take the shape of their container.
- _____ Molecules that can move (slide) over each other and have a definite size (volume), but not shape and can not be compressed.

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