

**Ch 5 Ions and Ionic Compounds Study Guide****Multiple Choice**

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. Which of the following electron configurations belongs to an element that is NOT chemically reactive?  
a.  $1s^2 2s^2 2p^3$  c.  $1s^2 2s^2 2p^5$   
b.  $1s^2 2s^2 2p^6 3s^2 3p^6$  d.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- \_\_\_\_\_ 2. In the compound sodium fluoride, NaF, the sodium atom loses one electron and the fluorine atom gains one electron to form ions that have electron configurations similar to  
a. helium. c. neon.  
b. oxygen. d. calcium.
- \_\_\_\_\_ 3. In many compounds, atoms of main-group elements form ions so that the number of electrons in the outermost energy levels of each ion is  
a. 2. c. 8.  
b. 6. d. 10.
- \_\_\_\_\_ 4. The electron configuration of nitrogen is  $1s^2 2s^2 2p^3$ . How many more electrons does nitrogen need to have an electron configuration similar to neon?  
a. 1 c. 5  
b. 3 d. 8
- \_\_\_\_\_ 5. Atoms of copper and iron  
a. generally form stable bonds with transition elements.  
b. have stable electron configurations.  
c. tend to form cations.  
d. tend to form anions.
- \_\_\_\_\_ 6. An anion  
a. is an ion with a negative charge.  
b. attracts ions with negative charges.  
c. results when an alkaline-earth metal loses one of its two outermost electrons.  
d. has more protons than electrons.
- \_\_\_\_\_ 7. When the octet rule is satisfied, the outermost \_\_\_\_\_ are filled.  
a. *d* and *f* orbitals c. *s* and *d* orbitals  
b. *s* and *p* orbitals d. *d* and *p* orbitals
- \_\_\_\_\_ 8. The elements of the \_\_\_\_\_ group satisfy the octet rule without forming compounds.  
a. halide c. alkali metal  
b. noble gas d. alkaline-earth metal
- \_\_\_\_\_ 9. Once an atom has full *s* and *p* orbitals in its outermost energy level,  
a. it is highly reactive only with alkali metals.  
b. it is highly reactive only with halogens.  
c. it can be combined with most elements.  
d. it has a stable octet and is unreactive.

# Ions And Ionic Compounds Study Guide

**M Walker**



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