energy level

CHAPTER 7 STUDY GUIDE

Ionic Compounds and Metals

Section 7.1 Ion Formation

chemical bond

In your textbook, read about chemical bonds and formation of ions.

electrons

Use each of the terms below just once to complete the passage.

nucleus	7.17.252.000		valence
nucleus	octet	pseudo-noble gas formations	valence
The force t	hat holds two atoms together i	s called a(n) (1)	
Such an attacl	ament may form by the attracti	on of the positively charged	
(2)	of one atom for the negatively charged		
	of another atom, or by the attraction of charged atoms, which are called		
	The attractions may also involve		
	electrons, which are the electrons in the outermost		
(6)	The (7)	are a family of elements	that have very
little tendency	to react. Most of these elemen	its have a set of eight outermost electrons	which is called a
stable (8)	. The rel	atively stable electron structures develope	ed by loss of
electrons in ex	ertain elements of groups 3, 4,	13, and 14 are called (9)	
For each stat	ement below, write true or fa-	lse.	
	 A positively charged ion is called an anion. 		
	 Elements in group 1 le charge. 	se their one valence electron, forming an	ion with a 1+
	12. Elements tend to react	so that they acquire the electron structure	e of a halogen.
	_ 13. A sodium atom tends t	o lose one electron when it reacts.	
	14. The electron structure formation.	of a zinc ion (Zn2+) is an example of a ps	eudo- noble gas
	_ 15. A Cl ⁺ ion is an example	e of a cation.	
	_ 16. The ending -ide is use	d to designate an anion.	
	 Nonmetals form a stab becoming anions. 	de outer electron configuration by losing	electrons and

noble gases

pous.

Ionic Compounds And Metals Study Guide

Brendan G. Carr

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