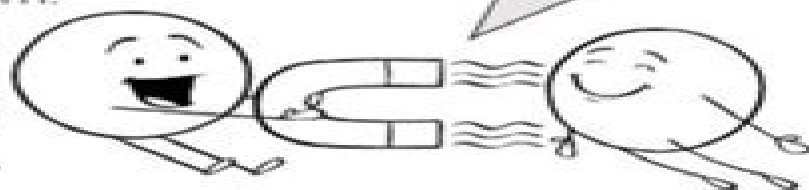


# CHEMICAL BONDS

## Ionic Bonding

Metals and non-metals react together to form \_\_\_\_\_ compounds. They are held together by ionic bonds. Ionic bonding is a force of \_\_\_\_\_ attraction between two \_\_\_\_\_ atoms (called \_\_\_\_\_). An atom becomes charged by either donating or accepting \_\_\_\_\_ electrons from another atom. A valence electron is one from the \_\_\_\_\_ shell of the atom.

Since opposites attract, there is an attraction between the positively charged and negatively charged ions that holds them together.

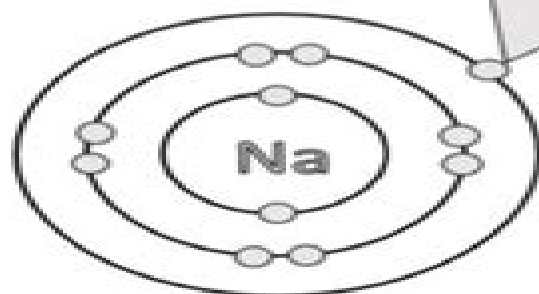


If an atom \_\_\_\_\_ an electron, it becomes positively charged. This is then called a \_\_\_\_\_. \_\_\_\_\_ always donate electrons. If an atom accepts an electron, it becomes \_\_\_\_\_. This is called an \_\_\_\_\_. Non-metals always accept electrons.

Atoms do this in order to become '\_\_\_\_\_'. A stable atom is one which has obtained a \_\_\_\_\_ of electrons.

*How do we know which atoms donate and which accept electrons?*

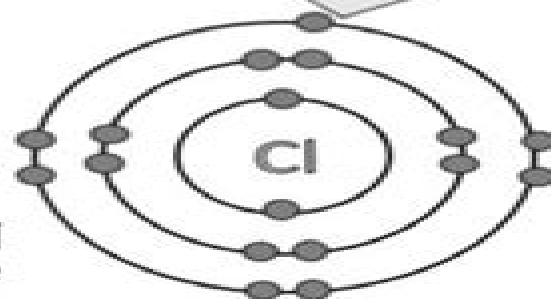
Sodium has \_\_\_\_\_ lone electron in the outer shell. It is easiest for this atom to \_\_\_\_\_ one electron than to try and find seven more to fill the outer shell.



All \_\_\_\_\_ donate electrons and therefore become cations.

All \_\_\_\_\_ accept electrons and therefore become anions.

Chlorine has \_\_\_\_\_ electrons in its outer shell. Therefore, it is easiest for Chlorine to \_\_\_\_\_ one electron than to \_\_\_\_\_ to try and donate seven.



# Ionic Bonds Worksheets For Middle School

**Camilla Rothe**



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## **Table of Contents Ionic Bonds Worksheets For Middle School**

1. Understanding the eBook Ionic Bonds Worksheets For Middle School
  - The Rise of Digital Reading Ionic Bonds Worksheets For Middle School
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Worksheets For Middle School
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic Bonds Worksheets For Middle School
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Worksheets For Middle School
  - Personalized Recommendations
  - Ionic Bonds Worksheets For Middle School User Reviews and Ratings
  - Ionic Bonds Worksheets For Middle School and Bestseller Lists
5. Accessing Ionic Bonds Worksheets For Middle School Free and Paid eBooks
  - Ionic Bonds Worksheets For Middle School Public Domain eBooks
  - Ionic Bonds Worksheets For Middle School eBook Subscription Services
  - Ionic Bonds Worksheets For Middle School Budget-Friendly Options
6. Navigating Ionic Bonds Worksheets For Middle School eBook Formats
  - ePub, PDF, MOBI, and More
  - Ionic Bonds Worksheets For Middle School Compatibility with Devices
  - Ionic Bonds Worksheets For Middle School Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ionic Bonds Worksheets For Middle School
  - Highlighting and Note-Taking Ionic Bonds Worksheets For Middle School
  - Interactive Elements Ionic Bonds Worksheets For Middle School
8. Staying Engaged with Ionic Bonds Worksheets For Middle School

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Ionic Bonds Worksheets For Middle School
- 9. Balancing eBooks and Physical Books Ionic Bonds Worksheets For Middle School
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ionic Bonds Worksheets For Middle School
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Bonds Worksheets For Middle School
  - Setting Reading Goals Ionic Bonds Worksheets For Middle School
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Bonds Worksheets For Middle School
  - Fact-Checking eBook Content of Ionic Bonds Worksheets For Middle School
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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