

IONIC

DEFINITION

an electrical attraction between a positively charged ion \oplus and a negatively charged ion \ominus



more protons \oplus
than electrons \ominus

atom loses
electron to
become \oplus
charge

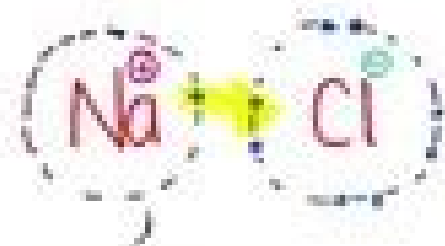


more electrons
than protons \oplus

atom gains
electron to
become
 \ominus charge

PROPERTIES

- high boiling point
- high melting point
- solid @ room temperature
- dissolve in water
- conduct electricity
- brittle



COVALENT

DEFINITION

chemical bond formed when electrons are shared.

single \longleftrightarrow share one pair of electrons

double \longleftrightarrow share two pairs of electrons

triple \longleftrightarrow shares three pairs of electrons

PROPERTIES

- do not conduct electricity
- lower melting point
- do not separate in water
- lower boiling point
- can be solid, liquid or gas at room temperature



METALLIC

DEFINITION

bond formed when many metal atoms share their pooled electrons

LIKE IONIC

give away
electrons

LIKE COVALENT

metals
share
electrons

PROPERTIES

- high boiling point
- high melting point
- good conductors of heat & electricity
- shiny



Ionic Bonding Notes Guide

Victor M. Corman



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Table of Contents Ionic Bonding Notes Guide

1. Understanding the eBook Ionic Bonding Notes Guide
 - The Rise of Digital Reading Ionic Bonding Notes Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonding Notes Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonding Notes Guide
 - User-Friendly Interface

4. Exploring eBook Recommendations from Ionic Bonding Notes Guide
 - Personalized Recommendations
 - Ionic Bonding Notes Guide User Reviews and Ratings
 - Ionic Bonding Notes Guide and Bestseller Lists
5. Accessing Ionic Bonding Notes Guide Free and Paid eBooks
 - Ionic Bonding Notes Guide Public Domain eBooks
 - Ionic Bonding Notes Guide eBook Subscription Services
 - Ionic Bonding Notes Guide Budget-Friendly Options
6. Navigating Ionic Bonding Notes Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bonding Notes Guide Compatibility with Devices
 - Ionic Bonding Notes Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonding Notes Guide
 - Highlighting and Note-Taking Ionic Bonding Notes Guide
 - Interactive Elements Ionic Bonding Notes Guide
8. Staying Engaged with Ionic Bonding Notes Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonding Notes Guide
9. Balancing eBooks and Physical Books Ionic Bonding Notes Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonding Notes Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Bonding Notes Guide
 - Setting Reading Goals Ionic Bonding Notes Guide
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Ionic Bonding Notes Guide
 - Fact-Checking eBook Content of Ionic Bonding Notes Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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