

20. Which of the following correctly represents the ionization of an atom?

- a) $\text{Cl(g)} + e^- \rightarrow \text{Cl}^-(\text{g})$
b) $\text{Na(g)} \rightarrow \text{Na}^+(\text{g}) + e^-$
 c) $\text{Na(s)} - e^- \rightarrow \text{Na}^+(\text{g})$
 d) $\text{Cl}_2(\text{g}) \rightarrow 2 \text{Cl(g)}$

21. Which of the following is likely to have the largest atomic radius?

- ~~a) H~~ ~~b) Mn~~ ~~c) Cl~~ **d) Rb** ~~e) Ag~~
 MOST LOW $\hat{=}$ LEFT

22. Which one of the following isoelectronic species has the smallest radius? *most protons*

- a) Mg^{2+} 12 p^+** d) F^- 9 p^+
 b) Na^+ 11 p^+ e) O^{2-} 8 p^+
 c) Ne 10 p^+

NOTE: explain your reasoning on the last page.

23. Which of the following has the greatest ionization energy? *smallest*

- a) K b) Ca c) Fe d) Ga **e) Br**
 all period 4

24. Which of the following has the lowest ionization energy? *LARGEST*

- a) Li b) Na c) K d) Rb **e) Cs**
 same family

25. The successive ionization energies for one of the period three elements is listed below. Which element is referred to?

| | |
|-------|---------------|
| E_1 | 577.4 kJ/mol |
| E_2 | 1,816 kJ/mol |
| E_3 | 2,744 kJ/mol |
| E_4 | 11,580 kJ/mol |
| E_5 | 15,030 kJ/mol |

} valence electrons
 } ~~core electrons~~

- a) Na ~~b) Mg~~ **c) Al** d) Si e) P

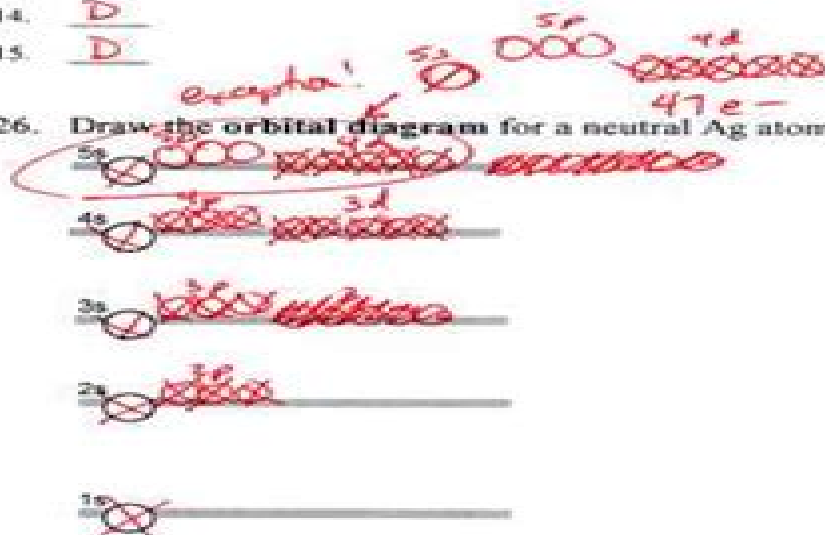
NOTE: explain your reasoning on the last page.

Multiple Choice Answers:

Use CAPITAL LETTERS for your answers....

- | | |
|--------------|---------------------------|
| 1. <u>B</u> | 16. <u>E</u> |
| 2. <u>D</u> | 17. <u>A</u> |
| 3. <u>C</u> | 18. <u>A</u> |
| 4. <u>D</u> | 19. <u>A</u> |
| 5. <u>C</u> | 20. <u>B</u> |
| 6. <u>E</u> | 21. <u>D</u> |
| 7. <u>A</u> | 22. <u>A</u> |
| 8. <u>C</u> | 23. <u>E</u> |
| 9. <u>B</u> | 24. <u>E</u> |
| 10. <u>B</u> | 25. A <u>C</u> |
| 11. <u>D</u> | |
| 12. <u>B</u> | |
| 13. <u>C</u> | |
| 14. <u>D</u> | |
| 15. <u>D</u> | |

26. Draw the orbital diagram for a neutral Ag atom: *excepta!* *47 e⁻*



Write the electron configuration for silver:

27. Long form:
 $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^1$
 28. Short form:
 $[\text{Kr}] 4d^{10} 5s^1$

Ionic And Metallic Bonding Answer Key

M Woodhall



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