- A <u>chemical bond</u> forms when 2 or more atoms rearrange <u>valence</u> <u>electrons</u> to increase <u>stability</u>.
- ionic bond forms when valence <u>clectrons</u> are <u>transferred</u>
 from one atom to another
 - cation atom <u>loses</u> electrons to become <u>+</u> charged
 - anion atom <u>sains</u> electrons to become ___ charged
 - In ionic compounds the ions are arranged in a <u>crystal</u> <u>lattice</u>

 Strong forces hold the ions together.

Properties of ionic compounds:

- high melting and boiling points
- hard not easily crushed
- conduct electricity when <u>melted</u> or

 dissolved because the ions are free to <u>move</u>
- covalent bond electrons are shared forming molecules
 - Covalent compounds have <u>weaker</u> forces holding the <u>molecules</u> together.

Properties of covalent compounds:

- lower melting and boiling points
- Many covalent compounds are <u>volatile</u> liquids or gases.
- soft ___ easier to __crush ______
- are not <u>conductors</u> of electricity
- electronegativity property that tells how strong an atom's <u>attraction</u> is for <u>electrons</u>
- Since oxygen has a <u>higher</u> electronegativity than hydrogen, oxygen holds onto shared electrons <u>tighter</u>, giving the oxygen a <u>partial</u> negative charge and the hydrogen a partial <u>positive</u> charge.

CHEMISTRY: A Study of Matter

Ionic Bonds And Notetaking Guide Answer

Emilie Sanchez

Ionic Bonds And Notetaking Guide Answer:

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Table of Contents Ionic Bonds And Notetaking Guide Answer

- 1. Understanding the eBook Ionic Bonds And Notetaking Guide Answer
 - The Rise of Digital Reading Ionic Bonds And Notetaking Guide Answer
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Ionic Bonds And Notetaking Guide Answer
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonds And Notetaking Guide Answer
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ionic Bonds And Notetaking Guide Answer

- Personalized Recommendations
- Ionic Bonds And Notetaking Guide Answer User Reviews and Ratings
- Ionic Bonds And Notetaking Guide Answer and Bestseller Lists
- 5. Accessing Ionic Bonds And Notetaking Guide Answer Free and Paid eBooks
 - Ionic Bonds And Notetaking Guide Answer Public Domain eBooks
 - Ionic Bonds And Notetaking Guide Answer eBook Subscription Services
 - Ionic Bonds And Notetaking Guide Answer Budget-Friendly Options
- 6. Navigating Ionic Bonds And Notetaking Guide Answer eBook Formats
 - o ePub, PDF, MOBI, and More
 - Ionic Bonds And Notetaking Guide Answer Compatibility with Devices
 - Ionic Bonds And Notetaking Guide Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonds And Notetaking Guide Answer
 - Highlighting and Note-Taking Ionic Bonds And Notetaking Guide Answer
 - Interactive Elements Ionic Bonds And Notetaking Guide Answer
- 8. Staying Engaged with Ionic Bonds And Notetaking Guide Answer
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonds And Notetaking Guide Answer
- 9. Balancing eBooks and Physical Books Ionic Bonds And Notetaking Guide Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonds And Notetaking Guide Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Bonds And Notetaking Guide Answer
 - Setting Reading Goals Ionic Bonds And Notetaking Guide Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Bonds And Notetaking Guide Answer

- Fact-Checking eBook Content of Ionic Bonds And Notetaking Guide Answer
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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