



2. What mass in grams of P_4O_6 could be made from 8.75 g P_4 and 12.50 g O_2 ?



$$8.75\text{g P}_4 \times \frac{1\text{mol P}_4}{123.88\text{g}} \times \frac{1\text{mol P}_4\text{O}_6}{1\text{mol P}_4} \times \frac{219.88\text{g}}{1\text{mol P}_4\text{O}_6} = \boxed{15.5\text{g P}_4\text{O}_6}$$

$$12.50\text{g O}_2 \times \frac{1\text{mol O}_2}{32.00\text{g}} \times \frac{1\text{mol P}_4\text{O}_6}{3\text{mol O}_2} \times \frac{219.88\text{g}}{1\text{mol P}_4\text{O}_6} = 28.63\text{g P}_4\text{O}_6$$

Flinn Scientific Pogil Activity Limiting And Excess Reactants Answers

Lifelique



Flinn Scientific Pogil Activity Limiting And Excess Reactants Answers:

Limiting Reactant and Percent Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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