Determining a Molar Volume

A Safer Experiment



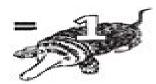
Introduction

The traditional method for calculating molar volumes involves generating oxygen gas—potassium chlorate is heated with numganese dioxide, a catalyst, to produce oxygen gas. This method is quite hazardous because large amounts of oxygen gas are produced and potassium oblorate is a powerful oxiditor of organic materials including the rubber stapper used in the ser-up. In fact, potassium chlorate is a frequent source of accidents on school premises. An easier and rater method presented in this laboratory activity for the calculation of molar volumes involves the use of carbon dioxide instead of oxygen gas.

Concepts

- 2/6/cite
- · Orlinbay violations
- · Talk at gas haw:





Materials

Dry ice, solid carbon disolde, CO₂, small piece about %' square

Balance

Baronseter

Extenuneper flash, Pyrex[®], 125-mL.

Ciral unted cylinder, 500-mL.
Rubber stopper, to fit Erlemmeyer flask.
Thermometer
Timps or insulated gloves (Zetex.*)

Safety Precautions

This activity requires the use of hazardous components and/or has the potential for hazardous reactions. Dry ice can cause freshite; never board dry ice directly with the shin; was appropriate invaluated gloves or use bugs whemever handling dry ice. Do not support the flack until dil of the dry ice has sublimed. Fadure to wast until all old prices are gone may cause the flack to shatter due to increased pressure. When chemical splack graggles, demical-resistant gloves, and a chemical-resistant apone. Please review current Material Safety Data Shorts for additional safety, handling, and disposal information.

Procedure

- Place the stopper in the Pyres Trienmeyer flask and mass to the nearest 0.01 g. Record this mass in the data table as the mass of the stoppered flask plus air. Recall that the flask is not actually empty—it contains air.
- Remove the stupper. Using tongs, or while wearing insulated gloves, place a small piece of dry ice in the Exhanneyer flask. Do not stupper the flask at this point.
- 3. Gently swirl the flask to aid in the sublimation of the dry ice.
- Immediately after the last bit of day ice has just sublimated and disappeared, rightly stopper the flask. Mass the flask and record this mass in the data table as the mass of the stoppered flask plus CO₂.
- Remove the stopper and fill the Erlemmeyer flack to the top lip with water. Stopper the flack again, Wipe any water off of the outside of the flack.
- Remove the stopper from the flask. Pour the water into a large graduated cylinder and measure the volume of water. The
 volume of water is equal to the volume of air (Step 1) or CO₂ (Step 4) in the flask. Record this volume in liters in the data
 table.
- 7. Record the room temperature in degrees Kelvin and the atmospheric pressure in atm in the data table.
- Record the density of air in the data table. The density of air is both temperature and pressure dependent. This value can be found in a chemistry textbook or reference manual if the temperature and pressure are known.
- 9. Calculate the mass of air in the flask by using the relationship that the density of air (Step 8) is equal to the mass of air

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