

Acids, Bases & Salts

Acids & Their Properties

An **acid** is a type of **sour** substance. When placed in water, acids release **hydrogen (H^+)** ions and negative ions, ions are negatively or positively charged atoms. When an acidic substance dissolves in water – hydrogen atoms break away from the acid molecules. This process forms hydrogen ions.



Hydrogen ions and chloride form when hydrochloric acid (HCl) is dissolved in water.



Hydrogen ion (H^+)

Characteristics of Acids

- have a **sour** taste
- react with metals
- turn **blue** litmus paper **red**
- conduct electric current



Indicators are substances that change color when they are added to acidic solutions or alkaline (basic) solutions.



Bases & Their Properties

A **base** is a type of **bitter** substance. When placed in water, bases release **hydroxide (OH^-)** ions into the water. Examples of a base include soap and baking soda.



The positive ions and hydroxide ions separate when sodium hydroxide (NaOH) is dissolved in water.



Hydroxide ion (OH^-)

Characteristics of Bases

- have a **bitter** taste
- feel **slippery**
- change **red** litmus paper to **blue**



bitter taste



slippery feel



Strength of Acids & Bases

Acids and bases vary in strength depending on the number of molecules that break apart to produce ions when dissolved in water. The more hydrogen ions (H^+) formed, the stronger the acid. Likewise, the more hydroxide ions (OH^-) formed the stronger the base.

The pH Scale

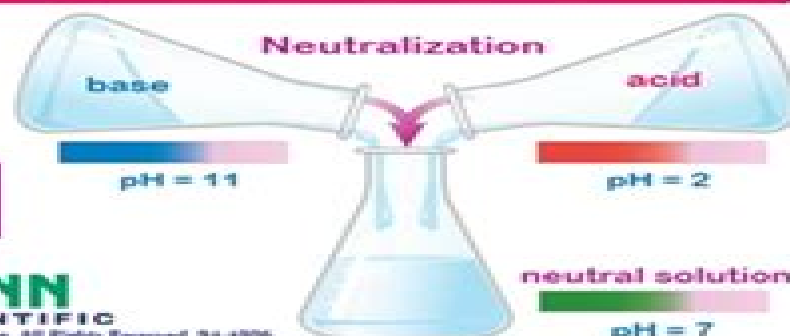
Scientists use a **pH scale** to measure the strength of acids and bases. The pH scale ranges from 0 to 14. Substances with pH of less than 7 are considered acids; the lower the pH, the stronger the acid. The pH of bases ranges from 7-14 – the higher the pH the stronger the base. Neutral solutions have a pH of exactly 7.



Salts

Acid-Base Reactions

A **neutralization** reaction occurs when an acid and a base react to form water and a salt. The hydrogen ions (H^+) from an acid react with the hydroxide ions (OH^-) from a base to form water.



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Flinn Scientific Acid Base Titrations Teacher Guide

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