

Stoichiometry and Solubility

Mole Ratios and Chemical Formulas

Introduction

Double replacement reactions are generally considered to be irreversible. The formation of an insoluble precipitate provides a driving force that makes the reaction proceed in one direction only. The purpose of this demonstration is to find the optimum mole ratio for the formation of a precipitate in a double replacement reaction and to use this information to predict the chemical formula of the precipitate.

Concepts

- Stoichiometry
- Mole ratios
- Double replacement reaction
- Solubility rules

Materials

Copper(II) chloride, CuCl_2 , 0.05 M, 250 mL
 Iron(III) nitrate, $\text{Fe}(\text{NO}_3)_3$, 0.1 M, 250 mL
 Sodium hydroxide, NaOH , 0.1 M, 250 mL
 Sodium phosphate, Na_3PO_4 , 0.05 M, 250 mL
 Graduated cylinders, 50-mL, 2
 Graduated cylinders, 100-mL, 7
 Stirring rods, large, 2

Safety Precautions

Copper(II) chloride, iron(III) nitrate, sodium hydroxide and sodium phosphate solutions are skin and eye irritants and are slightly toxic by ingestion. Avoid contact with eyes and skin. Wear chemical splash goggles, chemical-resistant gloves and a chemical-resistant apron. Please consult current Safety Data Sheets for additional safety information.

Procedure

Part A. Reaction of Iron(III) Nitrate with Sodium Hydroxide

- Label seven 100-mL graduated cylinders 1–7.
- Using a clean, 50-mL graduated cylinder, add the appropriate volume of iron(III) nitrate solution to each graduated cylinder as shown in Table 1.
- Use a second 50-mL graduated cylinder to add the appropriate volume of sodium hydroxide solution to each graduated cylinder, as shown in Table 1.

Table 1.

Cylinder	1	2	3	4	5	6	7
$\text{Fe}(\text{NO}_3)_3$, 0.1 M, mL	10	15	20	30	40	45	50
NaOH , 0.1 M, mL	50	45	40	30	20	15	10
$\text{Fe}^{3+}:\text{OH}^-$ Mole Ratio	1:5	1:3	1:2	1:1	2:1	3:1	5:1

- Use a large stirring rod to thoroughly mix the reactants. Observe the signs of chemical reaction in each cylinder. (Mixing the yellow-orange solution of iron(III) nitrate with the colorless sodium hydroxide solution gives a rust-colored precipitate and a pale yellow supernatant.)
- Let the reaction mixtures sit undisturbed for at least 10 minutes to allow the precipitates to settle. During this time, write the reactants on the board, and identify the possible products.
- After the precipitates have settled, record the volume of precipitate in each graduated cylinder.
- What mole ratio gave the maximum amount of precipitate? Explain.

Part B. Reaction of Copper(II) Chloride with Sodium Phosphate

- Label seven 100-mL graduated cylinders 1–7.
- Using a clean, 50-mL graduated cylinder, add the appropriate volume of copper(II) chloride solution to each graduated cylinder as shown in Table 2.
- Use a second 50-mL graduated cylinder to add the appropriate volume of sodium phosphate solution to each graduated cylinder as shown in Table 2.

Table 2.

Cylinder	1	2	3	4	5	6	7
CuCl_2 , 0.05 M, mL	10	20	25	30	35	40	50
Na_3PO_4 , 0.05 M, mL	50	40	35	30	25	10	10
$\text{Cu}^{2+}:\text{PO}_4^{3-}$ Mole Ratio	1:5	1:2	2:3	1:1	3:2	2:1	5:1

Flinn Chemtopic Labs Mole Ratios Key

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