

6) Use the following information to determine the heat of reaction for $\text{NO}_{(g)} + \frac{1}{2} \text{O}_{2(g)} \rightarrow \text{NO}_{2(g)}$



7) Calculate the molar heat of formation for copper (I) oxide using the following information:



8) Use heats of formation from the table in the appendix of your textbook to calculate the enthalpy of reaction for the following:



9) Review: Calculate the amount of heat released when 4.79 g of ethene (C_2H_4) is burned in excess oxygen. ($\Delta H_{\text{combustion}} = -1.39 \times 10^3 \text{ kJ/mol}$). Write the balance thermochemical equation for the reaction.

10) Review: An ice cube with a mass of 40.0 g melts in water originally at 25°C . How much heat is absorbed by the melting ice cube from the surrounding water?

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