

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H}$$

C₃H₆

2. A compound is found to contain 50.0 g of sulfur, 43.7 g of oxygen, and 6.27 g of hydrogen. Find the empirical formula.

$$50.0 \text{ g S} \times \frac{1 \text{ mol S}}{32.07 \text{ g S}} = 1.56 \text{ mol S}$$

$$43.7 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 2.73 \text{ mol O}$$



$$6.27 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.21 \text{ mol H} \quad \left(\frac{6.21 \text{ mol H}}{1.56 \text{ mol S}} = 4 \right)$$

3. A 10.0 g sample of a compound contains 2.0 g of carbon, 1.0 g of hydrogen, and 7.0 g of oxygen. Find the empirical formula.

$$2.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.17 \text{ mol C}$$



$$1.0 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 0.99 \text{ mol H} \quad \left(\frac{0.99 \text{ mol H}}{0.17 \text{ mol C}} = 6 \right)$$

$$7.0 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.44 \text{ mol O} \quad \left(\frac{0.44 \text{ mol O}}{0.17 \text{ mol C}} = 3 \right)$$

4. A 25.00 g sample of an unknown compound contains 12.50 g carbon, 12.50 g hydrogen, and 0.00 g oxygen. Find the empirical formula.

$$12.5 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.04 \text{ mol C}$$



$$12.5 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 12.4 \text{ mol H}$$

$$0.00 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.00 \text{ mol O} \quad \left(\frac{12.4 \text{ mol H}}{1.04 \text{ mol C}} = 12 \right)$$

Empirical Formula Pogil Answers

Jessica J Manson



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