

Example 1: Calculate the empirical formula for a substance containing 1.11 g of Se and 1.61 g of F.

$$1.11 \text{ g } \cancel{\text{Se}} \times \frac{1 \text{ mol Se}}{78.971 \text{ g } \cancel{\text{Se}}} = 0.0141 \text{ mol Se}$$

$$1.61 \text{ g } \cancel{\text{F}} \times \frac{1 \text{ mol F}}{18.998 \text{ g } \cancel{\text{F}}} = 0.0847 \text{ mol F}$$

$$\text{Se} \frac{0.0141}{0.0141} \text{F} \frac{0.0847}{0.0141} = \text{SeF}_6$$

Empirical Formula Answers Only

**Donald J. Wink, Sharon Fetzer-
Gislason, Sheila McNicholas**



Empirical Formula Answers Only:

Empirical Formula Answers Only Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has become more evident than ever. They have the ability to inspire, provoke, and ignite change. Such is the essence of the book **Empirical Formula Answers Only**, a literary masterpiece that delves deep into the significance of words and their effect on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall impact on readers.

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