EMPIRICAL AND MOLECULAR FORMULAE

EMRIRICAL FORMULA - This represents the simplest whole number ratio of atoms within a molecule

MOLECULAR FORMULA - This represents the exact numbers of each atom in a molecule

Empirical formulae can be calculated from either % mass composition data or mass composition data. Molecular formula can then be calculated if the molecular mass of the compound is known.

EXAMPLE 1 - from % composition

A substance containing carbon and hydrogen and oxygen only contains 73.47% carbon and 10.20% hydrogen by mass, its relative molecular mass is 98. Calculate its molecular formula.

Step 1 calculate empirical formula

ATOMS	CARBON	HYDROGEN	OXYGEN
PERCENTAGE MASS	73.47	10.20	(100 - (73.47 + 10.20) = 16.33
ATOMIC MASS	12	1	16
MASS/ATOMIC MASS	6.1225	10.20	1.020
RATIO(divide by smallest)	6	10	1

Empirical formula = C.H.O

Step 2 use Mo to find molecular formula

 $M_{\rm E}$ = mass ($C_6H_{10}O$) \times n

98 = (98) x n n = 1

molecular formula =

CHO

EXAMPLE 2 - from mass data

A substance contains 12.28g nitrogen, 3.51g hydrogen, 28.07g sulphur and 56.10g oxygen, its Me is 228 .Calculate its molecular formula.

Empirical And Molecular Answers And Questions

Lei Shi

Empirical And Molecular Answers And Questions:

Empirical And Molecular Answers And Questions Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Empirical And Molecular Answers And Questions**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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