

# Empirical and Molecular Formulas Worksheet

Key

## Objectives:

- be able to calculate empirical and molecular formulas

## Empirical Formula

1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l}
 \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} \approx 1.96 \quad 2 \\
 \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} \approx 5.88 \quad 6 \\
 \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1
 \end{array}$$

$\text{C}_2\text{H}_6\text{O}$

2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{array}{l}
 \text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5} \\
 \text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5
 \end{array}$$

$\text{Au}_2\text{O}_3$

3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\begin{array}{l}
 \text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad \cdot \quad 2 \\
 \text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad \cdot \quad 8 \\
 \text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad \cdot \quad 1 \\
 \text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad \cdot \quad 4
 \end{array}$$

$\text{N}_2\text{H}_8\text{S}_4\text{O}_4$

## Molecular Formula

4) Empirical formula of a substance is  $\text{CH}_2\text{O}$ . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 0 + 16) = 30$$

$$\frac{180}{30} = 6 \quad \text{C}_6\text{H}_{12}\text{O}_6$$

5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\begin{array}{l}
 \text{C} \quad \frac{1.388}{3.585} \rightarrow 38.7\% \rightarrow \frac{3.87\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1 \\
 \text{H} \quad \frac{0.345}{3.585} \rightarrow \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3 \\
 \text{O} \quad \frac{1.850}{3.585} \rightarrow \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1
 \end{array}$$

$\text{C}_2\text{H}_6\text{O}_2$

$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 0 + 16)$

$$\frac{62}{31.03} = 2$$

1	$\text{C}_2\text{H}_6\text{O}$
2	$\text{Au}_2\text{O}_3$
3	$\text{NH}_4\text{SO}_4$
4	$\text{C}_2\text{H}_6\text{O}_2$
5	$\text{C}_6\text{H}_{12}\text{O}_6$

# Empirical Molecular Formula Answers

**IM Harris**



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