

Let the relative atomic mass of element Z be z .

Element	Z	Bromine, Br
Mass (g)	2.07	$3.67 - 2.07 = 1.60$
Number of moles of atoms	$\frac{2.07}{z}$	$\frac{1.60}{80} = 0.02$

Based on the empirical formula ZBr_2 , the ratio of atoms of Z : Br is 1 : 2.

Therefore, $\frac{2.07}{z} : 0.02 = 1 : 2$

$$\frac{2.07}{z} \div 0.02 = 1 \div 2$$

$$\frac{2.07}{0.02z} = \frac{1}{2}$$

$$z = \frac{2 \times 2.07}{0.02}$$

$$= 207$$

The relative atomic mass of element Z is 207.

Empirical Formula For Copper Oxide

Michael Brown

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Empirical Formula For Copper Oxide:

Reviewing **Empirical Formula For Copper Oxide**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Empirical Formula For Copper Oxide**," an enthralling opus penned by a highly acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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