Molarity = Moles of solute / Liters of Solution (abbreviation = M)
Molality = Moles of solute / Kg of Solvent (abbreviation = m)
Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)
Mass Percent = mass of solute / mass of solution

How many moles of ethyl alcohol, C₂H₂OH, are present in 65 mL of a 1.5 M solution?

65 mL C ₂ H ₅ OH	1 L C ₂ H ₅ OH	1.5 mole C ₂ H ₅ OH	= 0.098 moles C ₂ H ₅ OH
	1000 mL C ₂ H ₅ OH	1 L C ₂ H ₂ OH	

2. How many liters of a 6.0 M solution of acetic acid CH₂COOH, contain 0.0030 moles of acetic acid?

0.0030 moles CH ₃ COOH	1 L solution	-	5.0 x 104 L solution
	6.0 moles CH ₃ COOH		

3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?

85 g KOH	1 mole	1 kg solvent	1000 g solvent	= 5.0 x 102 g solvent
	56.11 g KOH	3.0 mole KOH	1 kg solvent	

5.0 x 102 g solvent + 85 g KOH = 590 g solution

4. If you dissolve 0.70 moles of HCI in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?

2.5 L solution

A solution is prepared by adding 2.0 L of 6.0 M HCI to 500 mL of a 9.0 M HCI solution. What is the molarity of the new solution? (Remember, the volumes are additive)

2.0 L solution	6.0 moles HCI	= 12 moles HCI
	1 L solution	
0.500 L solution	9.0 moles HCI	= 4.5 moles HCI
	1 L solution	
12 moles HCI +	4.5 moles HCI =	17 moles
2.0 L solution +	0.500 L solution	= 2.5 L solution
M = 17 moles	= 6.8 M H	CI

6. Convert the following Molarities to Normalities.

a.	2.5 M HCI	_	2.5	- N
b.	1.4 M H ₂ SO ₄	-	2,8	24

- c. 1.0 M NaOH = 1.0 N
- d. 0.5 M Ca(OH)₂ = 1 N

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