

Qualitative analysis of an organic compound showed that it contained only carbon, hydrogen and oxygen. A quantitative study of the same compound was performed, in which a 0.5096 g sample was burnt in excess oxygen to produce 0.4160 g of water and 700.7 mL of carbon dioxide, collected at 100.0 °C and 102.8 kPa.

373K

(a) Determine the empirical formula of the compound.

(10 marks)

C	H	O
<p>0.7007 L of CO_2</p> <p>$n(\text{C}) = n(\text{CO}_2)$</p> <p>$PV = nRT \rightarrow n = \frac{PV}{RT}$</p> <p>$n(\text{CO}_2) = \frac{102.8 \times 0.7007}{8.315 \times 373}$</p> <p>$= 2.3218 \times 10^{-2} \text{ mol}$</p> <p>$n(\text{C}) = nM = 0.27885 \text{ g}$</p>	<p>0.416 g of H_2O</p> <p>$n = \frac{m}{M} = \frac{0.416}{18.016} = 2.3091 \times 10^{-2} \text{ mol}$</p>	<p>$m(\text{O}) = 0.5096 - m(\text{C}) - m(\text{H})$</p> <p>$= 0.18423 \text{ g}$</p> <p>$n = \frac{m}{M} = \frac{0.18423}{16} = 1.1514 \times 10^{-2}$</p>

EMPIRICAL FORMULA

EXAM QUESTION

$$\frac{2.3218}{1.1514} = 2$$

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The Princeton Review



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