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they are the same as basic solutions

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nh₂ and s²⁻ are ch₄ nh₃ and hs⁻ respectively next we consider that acidity increases with positive charge on the molecule

thus ruling out that s²⁻ is the weakest base

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compounds have characteristics of both acids and bases in the same molecule we will explore the relationship between

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molecule that gains a hydrogen ion becomes this which is positively charged what did lewis propose an acid accepts a pair of

electrons during a reaction while a base donates a pair of electrons

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