QUESTION 8

8.1 8,78 g of oxygen gas is required to react with an unknown mass of hydrogen gas to form water vapour. The balanced equation for this is:

$$O_{2(g)} + 2H_{2(g)} \rightarrow 2H_2O_{(g)}$$

- 8.1.1 Define the term exothermic reaction. (2)
- 8.1.2 Is this reaction a synthesis or a decomposition reaction?

 Give a reason for the answer. (2)
- 8.1.3 Calculate the number of moles of water vapour formed. (2)
- 8.2 The reaction between zinc and dilute hydrochloric acid is represented by the balanced equation below:

$$Zn(s) + 2HC\ell(aq) \rightarrow ZnC\ell_{+}(aq) + H_{+}(g)$$

The concentration of hydrochloric acid solution is 0,1 mol.dm⁻³. The mass of zinc chloride formed is 0,85 g.

- 8.2.1 Define the term concentration in words. (2)
- 8.2.2 Is the above reaction an example of a redox reaction? Write down YES or NO and give a reason for the answer. (2)
- 8.2.3 Calculate the volume of hydrochloric acid solution needed to react with the zinc to form 0,85 g of ZnCl₂. (5)

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