



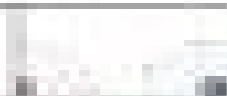
Name: _____ Date: _____

Student Exploration: Ionic Bonds

Readings for
Now

Read the background:

- [Ionic Bonds](#)
- [Electron-dot diagram conventions and ionic bonding](#)



Background: Some of the particles that make up matter have an electrical charge. Electrons are negatively charged, while protons are positively charged. Particles with opposite charges (+ and -) are **attracted** toward the other charges (+ and + or - and -) **repel**.

Question: What happens when atoms gain or lose electrons?

1. **Observe:** Electrons move around the nucleus of atoms in specific shells, known as **energy levels**. Electrons outside the nucleus are **valence electrons**. How many valence electrons does the element **fluorine** have? How many valence electrons does the element **potassium** have?

2. Observe the results and clicking arrows, assuming that fluorine has one valence electron, how many electrons did the potassium atom have?

Fluorine: _____ Potassium: _____

3. How many electrons did you add to the potassium atom to make it have the same number of valence electrons as fluorine? How many electrons did it need to **lose**?

Fluorine: _____ Potassium: _____

4. **Observe:** What happens when you click on a valence electron of one atom and drag it to the other? What happens? How many valence electrons did the fluorine atom have?

Fluorine: _____ Potassium: _____

5. **Observe:** What happens when you click on a valence electron of one atom and drag it to the other? What happens? How many valence electrons did the potassium atom have?

6. What does the charge of each atom have? Fluorine: _____ Potassium: _____

The **net ionic charge** is the sum of the charges of the ions. What is the net ionic charge?

7. How many electrons did you add to the potassium atom to make it have the same number of valence electrons as fluorine?

Fluorine: _____ Potassium: _____



Get The Gizmo Ready Activity A Ions

Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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