

Calorimetry Lab Report

Lab Members: Dylan Craig, and Brayden

Objective: The objective of this lab is to find out if a fishing weight was actually made of lead, to investigate the first law of thermodynamics and Calorimetry, and to design an experiment by first creating an analytical model as a guide.

Introduction: By combining the first law of thermodynamics with Calorimetry, one can determine if a fishing weight that claims to be lead, is actually made of lead or not. The first law of thermodynamics states that you can change the internal energy of a system by doing work or applying heat to a system.

$$\Delta E_{\text{system}} = W_{\text{on}} + Q$$

This can help determine the heat capacity of the lead weight. The heat [Q] is a physical quantity that characterizes the transfer of energy in the form of heat from an environment to a system (pg. 1, Calorimetry).

$$Q = mc \Delta T$$

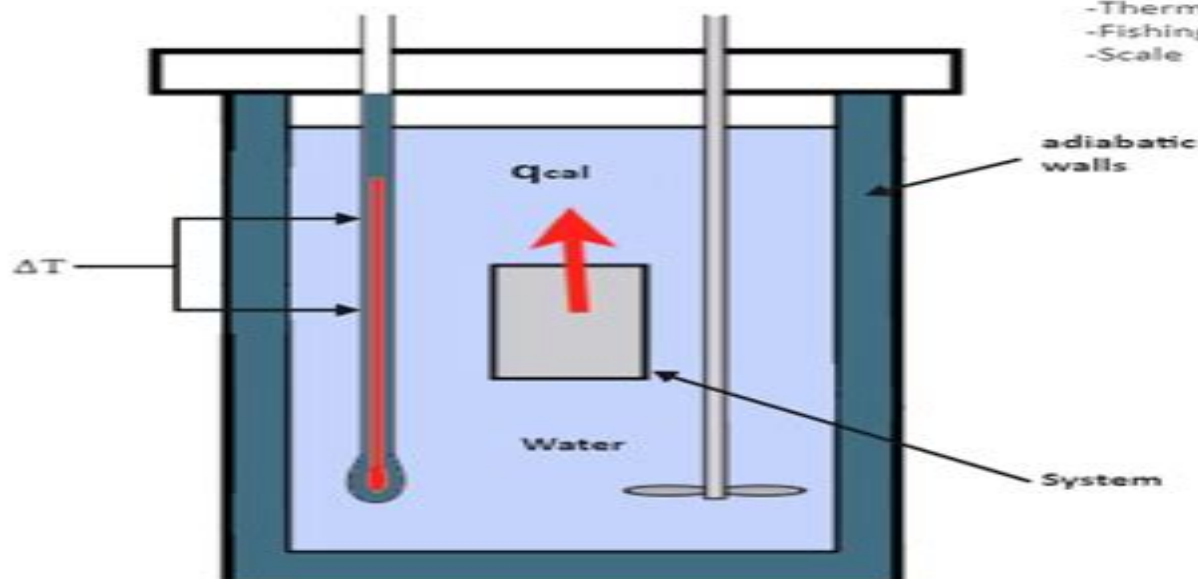
It is important to note that when one is analyzing two objects that are exchanging heat, the conservation of energy is:

$$Q_{\text{hot to cold}} + Q_{\text{cold to hot}} = 0$$

Experimental Setup:

Equipment Used:

- Calorimeter
- Water
- Thermometer
- Fishing Weight
- Scale



Enthalpy Calorimeter Lab Report

AN Whitehead



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of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment s Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

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