

Relevantly: Maria Fum will get:

Name: _____ Date: _____

Others will work and complete your final answer:

1. What is the pH of a 0.00100 M HCl solution?

$$\text{pH} = -\log[\text{H}^+] = -\log[1.0 \times 10^{-3}] = 3.00 \quad \boxed{3.00}$$

2. The molar mass of HCl is 36.46 g/mol. What is the pH of a 0.00100 M HCl solution?

$$\text{pH} = -\log[\text{H}^+] = -\log[1.0 \times 10^{-3}] = 3.00 \quad \boxed{3.00}$$

3. The pH of a 0.00100 M HCl solution is _____.

$$\text{pH} = -\log[\text{H}^+] = -\log[1.0 \times 10^{-3}] = 3.00 \quad \boxed{3.00}$$

$$\text{pH} = -\log[\text{H}^+] = -\log[1.0 \times 10^{-3}] = 3.00 \quad \boxed{3.00}$$

4. Human blood has a pH of 7.4. Calculate both the $[\text{H}^+]$ and $[\text{OH}^-]$ concentrations.

$$\text{pH} = -\log[\text{H}^+] \quad \Rightarrow \quad 7.4 = -\log[\text{H}^+] \quad \Rightarrow \quad 1.0 \times 10^{-7} \quad \boxed{1.0 \times 10^{-7} \text{ M}}$$

$$[\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \quad \Rightarrow \quad (1.0 \times 10^{-7})([\text{OH}^-]) = 1.0 \times 10^{-14} \quad \boxed{[\text{OH}^-] = 1.0 \times 10^{-7} \text{ M}}$$

5. The greater the $[\text{H}^+]$ concentration, the stronger _____ the pH.

6. An aqueous solution has a pH of 10.2. What has been its pH? 10.2. Which solution has the stronger concentration of $[\text{OH}^-]$ ions? The 10.2 solution has a higher concentration of $[\text{OH}^-]$ ions.

7. If a solution is basic, _____, then $[\text{H}^+]$ concentration equals the $[\text{OH}^-]$ concentration.

More Fun With Ph Answer Key

Michael Rajnik



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