



Kinetic Molecular Theory

Postulates of the Kinetic Molecular Theory of Gases

1. Gases consist of tiny particles (atoms or molecules) These particles are so small, compared with the distances between them, that the volume (size) of the individual particles can be assumed to be negligible (zero).
2. The particles are in constant random motion, colliding with the walls of the container. These collisions with the walls cause the pressure exerted by the gas.
3. Molecular collisions are perfectly elastic. This means that when the molecules collide with each other, and the walls of the container, we assume they lose no energy at all – they never slow down, and always stay at the same speed!
4. The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas – that is, the higher the temperature, the higher the energy of the gas!

Kinetic Molecular Theory Pogil Answer

Padhraic Smyth



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