

8. 44. The conditions used in an electrolytic cell can determine the products formed.

44. Draw an electrolytic cell showing the electrolysis of molten sodium chloride. Include in the diagram the direction of the electron flow, the nature of electrodes and state the half-reactions for the products formed at each electrode. (11)

Blank box for drawing the electrolytic cell.

45. Deduce the equations for the formation of the major product at the positive electrode (anode) when the following aqueous solutions are electrolysed.

- dilute sodium chloride
- concentrated sodium chloride

Blank box for writing equations.

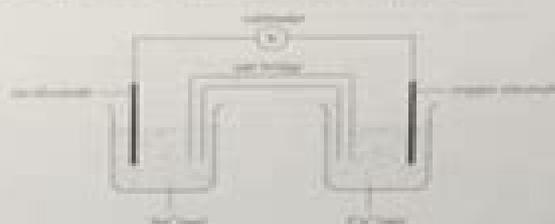
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Time used

Question Continued

46. A silver cell constructed from two half-cells is illustrated below.



46. Use Table 14 of the Data Booklet to deduce the equation for the spontaneous reaction occurring in this cell. (11)

Blank box for writing the equation for the spontaneous reaction.

47. Write the standard potential for this cell. (1)

Blank box for writing the standard potential.

48. State the conditions necessary for the operation of this cell to equal that indicated in part 46(a) using the Data Book, Table 14. (1)

Blank box for stating the conditions necessary for the operation of the cell.

If the question continues on the following page



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