

8. (a) The conditions used in the electrolysis will now determine the products formed.

(i) Draw an electrolytic cell showing the electrodes or vessels suitable for brackish water. Include in the diagram the direction of the electron flow, the polarity of electrodes and write the half-reactions for the products formed at each electrode. [11]

(ii) Deduce the equations for the formation of the major product at the positive electrode (anode) when the following aqueous solutions are electrolysed.

- dilute sodium chloride
- concentrated sodium chloride

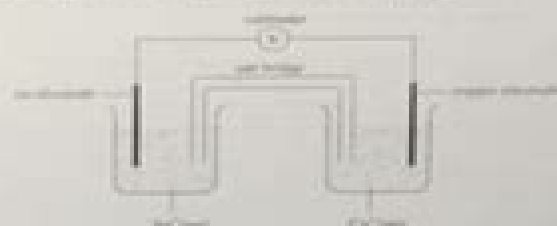
(The question continues on the following page)



Have over

(Question 8 continued)

- (b) A salted cell is constructed from two half-cells as illustrated below.



(i) Use Table 14 of the Data Booklet to deduce the equation for the spontaneous reaction occurring in this cell. [10]

(ii) Write the standard potential for this cell. [10]

(iii) State the conditions necessary for the potential of the cell to equal that calculated in part (ii) using the data from Table 14. [10]

(The question continues on the following page)



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