Study Guide for Chemistry EOC Assessment

Your test will consist of multiple-choice questions over the following topics.

- 1. Metric units for mass, length, volume
- Metric to metric conversions (K H Da m/l/g D C M)
- Know how to put numbers in scientific notation and take them out.
- 4. Significant Digit rules
- Difference between chemical and physical properties and changes. There will be multiple questions on this.
- 6. Ways to tell if a chemical change (chemical reaction) is taking place.
- 7. Pure substances and mixtures
 - a. heterogeneous mixture
 - b. homogeneous mixture
- 8. Density & how it relates to floating and sinking
- Use the equation for density
- Density = Mass/volume (g/ml) (g) (ml)
- Know the phases of matter and the characteristics of each phase.
- 11. Law of Conservation of Matter
- 12. What is a mole in chemistry?
- 13. Convert grams to moles, and moles to grams
- Define an atom
- 15. Structure of an atom and where each subatomic particle is located.
 - a. Nucleus (contains protons and neutrons)
 - b. Protons (positive)
 - c. Neutrons (neutral)
 - d. Electrons (negative)
- 16. Isotopes
- 17. Elements
- 18. If given atomic # and mass, find the number of protons and neutrons in an atom.
- 19. Find the number of valence electrons, and tell what it says about reactivity.
- 20. Periodic Table
 - a. Know about the families
 - b. Group vs. period
- 21. Know about the properties of metals and nonmetals.
 - a. Ductile, malleable, conductors
- 22. Ionic & Covalent Compounds

How do they form? (Transfer or gain electrons) Characteristics of Each

- Be able to name ionic compounds
- 24. Predict oxidation numbers
- 25. Chemical Equations:

Know the parts of an equation

5 main types of reactions-know about them and be able to identify them (Synthesis, Decomposition, Single Displacement, Double Displacement, Combustion)

26. Acids & Bases-

Macro and Microscopic properties of each

What ions do they form, how do they affect litmus, what is their pH range?

Identify examples of each- Mono, di, and polyprotic acids

Tell if they are strong or weak

Acid-base indicator

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Judd E. Hollander

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