

Chapter Assessment Chemistry: Matter and Change • Chapter 4

Reviewing Vocabulary

Match each definition in Column A with the term in Column B.

Column A

Column B

1. Radiation deflected toward the positively charged plate
 2. Atoms with the same number of protons but different numbers of neutrons
 3. High-energy radiation that has no charge and no mass
 4. The smallest particle of an element that retains the properties of that element
 5. The weighted average mass of an element's isotopes
 6. The center-most part of an atom where the protons and neutrons are contained
 7. Radiation deflected toward the negatively charged plate
 8. The rays and particles emitted by radioactive material
 9. Equal to $1/12$ the mass of a carbon-12 atom
 10. Ray of radiation traveling from the cathode to the anode
 11. Process not requiring energy by which unstable nuclei lose energy
 12. States that all matter is composed of atoms
 13. Process whereby some substances spontaneously emit radiation
- Compare and contrast each pair of related terms.
14. mass number, atomic number
 15. nuclear reaction, nuclear equation

- a. atom
- b. nucleus
- c. atomic mass
- d.
- e. gamma ray
- f. alpha radiation
- g. beta
- h. atomic
- i.
- j. Dalton's atomic theory
- k. cathode ray
- l.
- m.

Understanding Main Ideas

Use the periodic table to identify each element described below.

1. atomic number 63
2. 18 protons
3. 44 protons and 44 electrons
4. atomic number 24
5. 21 protons
6. atomic number 55

In the space at the left, write true if the statement is true; if the statement is false,

change the italicized term to make it true.

7. An atom's nucleus contains its protons and electrons.
8. Neutrons have no electrical charge.
9. Beta particles have a charge of 2^+ .
10. An alpha particle consists of two protons and two electrons.

Complete the table below.

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E Durkheim



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