#### Stoichiometry Study Guide

#### Students should be able to...

- Explain mole ratio
- · Convert moles, grams, atoms/molecules
- Calculate theoretical yields when given the limiting reactant (Convert: moles A to moles B, grams A to moles B, moles A to grams B, grams A to grams B)
- Explain limiting reactant and excess.
- Calculate theoretical yields by finding the limiting reactant
- Calculate percent yield.
- Calculate percent composition.
- Calculate theoretical yield, limiting reactant, excess, and percent yield for a doublereplacement reaction (laboratory)

**Practice Problems** 

Find the molar mass of each compound.

- 1. Bala
- 2. AgC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>

Convert.

- 3. 0.4 moles of Sn to atoms
- 4. 32 g of K<sub>2</sub>O to molecules
- 5. 3.7 g of H<sub>2</sub> to moles
- 6.800 g of CaCl<sub>2</sub> to moles
- 7. 0.75 moles of H<sub>2</sub>O to grams
- 8. 9.0 q of Cl<sub>2</sub> to atoms
- 9. 0.5 moles of KNO<sub>3</sub> to grams

Find the percent composition for each element in the compound.

- 10. Min<sub>2</sub>O<sub>7</sub>
- 11. AUNOs

Stoichiometry

12. MgO + H<sub>2</sub>O ----> Mg(OH)<sub>2</sub>

If you started with 3.7 moles of magnesium oxide, how many moles of magnesium hydroxide would be produced?

13. Na + I<sub>2</sub> ---> Nal

If 2.8 grams of sodium were used for the reaction, and an excess of iodine was present, how many moles of sodium iodide would be produced?

14. No + 3 Ho ----> 2 NHo

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**M** Mosston

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