

## Chapter 5 Assessment

pages 166–169

### Section 5.1

#### Mastering Concepts

**34.** Define the following terms.

**a.** frequency

Frequency is the number of waves that pass a given point per second.

**b.** wavelength

Wavelength is the shortest distance between equivalent points on a continuous wave.

**c.** quantum

A quantum is the minimum amount of energy that can be lost or gained by an atom.

**d.** ground state

An atom's ground state is its lowest allowable energy state.

**35.** Arrange the following types of electromagnetic radiation in order of increasing wavelength.

**a.** ultraviolet light

**b.** microwaves

**c.** radio waves

**d.** X rays

**d.** X rays, **a.** ultraviolet light, **b.** microwaves, **c.** radio waves

**36.** A gamma ray has a frequency of  $2.88 \times 10^{21}$  Hz. What does that mean?

$2.88 \times 10^{21}$  gamma-ray electromagnetic waves of this frequency pass a given point per second.

**37.** What is the photoelectric effect?

a phenomenon in which a metal emits electrons when light of a sufficient frequency shines on it

**38. Neon Sign** How does light emitted from a neon sign differ from sunlight?

The light from a neon sign contains only certain visible colors, while sunlight contains the full spectrum of colors.

**39.** Explain Planck's quantum concept as it relates to energy lost or gained by matter.

According to Planck, for a given frequency,  $\nu$ , matter can emit or absorb energy only in discrete quanta that are whole-number multiples of  $h\nu$ , where  $h$  is Planck's constant.

**40.** How did Einstein explain the photoelectric effect?

He proposed that photons must have a certain minimum energy level, or threshold, value to cause the ejection of a photoelectron.

**41. Rainbow** What are two differences between the red and green electromagnetic waves in a rainbow?

The red waves have a longer wavelength and a lower frequency.

**42. Temperature** What happens to the light emitted by a heated, glowing object as its temperature increases?

The color of the light changes as the object acquires more and more energy.

**43.** What are three deficiencies of the wave model of light related to light's interaction with matter?

The wave model does not explain the photoelectric effect, atomic emission spectra, and why matter emits different frequencies of light at different temperatures.

**44.** How are radio waves and ultraviolet waves similar? How are they different?

Both types of waves travel at the same speed in a vacuum,  $3.00 \times 10^8$  m/s. Radio waves have a longer wavelength and lower frequency than ultraviolet waves.

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## **Table of Contents Glencoe Chemistry Chapter 5 Assessment**

1. Understanding the eBook Glencoe Chemistry Chapter 5 Assessment
  - The Rise of Digital Reading Glencoe Chemistry Chapter 5 Assessment
  - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Chemistry Chapter 5 Assessment
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Glencoe Chemistry Chapter 5 Assessment
  - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Chemistry Chapter 5 Assessment
  - Personalized Recommendations
  - Glencoe Chemistry Chapter 5 Assessment User Reviews and Ratings

- Glencoe Chemistry Chapter 5 Assessment and Bestseller Lists
- 5. Accessing Glencoe Chemistry Chapter 5 Assessment Free and Paid eBooks
  - Glencoe Chemistry Chapter 5 Assessment Public Domain eBooks
  - Glencoe Chemistry Chapter 5 Assessment eBook Subscription Services
  - Glencoe Chemistry Chapter 5 Assessment Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - Glencoe Chemistry Chapter 5 Assessment Compatibility with Devices
  - Glencoe Chemistry Chapter 5 Assessment Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Glencoe Chemistry Chapter 5 Assessment
  - Highlighting and Note-Taking Glencoe Chemistry Chapter 5 Assessment
  - Interactive Elements Glencoe Chemistry Chapter 5 Assessment
- 8. Staying Engaged with Glencoe Chemistry Chapter 5 Assessment
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Glencoe Chemistry Chapter 5 Assessment
- 9. Balancing eBooks and Physical Books Glencoe Chemistry Chapter 5 Assessment
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Glencoe Chemistry Chapter 5 Assessment
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Glencoe Chemistry Chapter 5 Assessment
  - Setting Reading Goals Glencoe Chemistry Chapter 5 Assessment
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Glencoe Chemistry Chapter 5 Assessment
  - Fact-Checking eBook Content of Glencoe Chemistry Chapter 5 Assessment
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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